ENGAGE

Why does the atomic number but not the mass number change in positron emission?

SAMPLE PROBLEM 5.4 Writing a Nuclear Equation for Positron Emission

TRY IT FIRST

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

SOLUTION GUIDE

ANALYZE THE PROBLEM	Given	Need	Connect
	Mn-49, positron	balanced nuclear	mass number, atomic
	emission	equation	number of new nucleus

STEP 1 Write the incomplete nuclear equation.

$$^{49}_{29}Mm \longrightarrow 7 + ^{9}_{47}e$$

STEP 2 Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = 7 + 0$$

$$49 - 0 = ?$$

$$49 - 0 = 49$$
 (mass number of new nucleus)

STEP 3 Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = 7 + 1$$

$$25 - 1 = 7$$

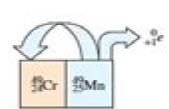
STEP 4 Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written 50r.

SIEP 5 Complete the nuclear equation.

STUDY CHECK S.4.

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

ANSWER



<u>Introduction To Chemical Answers Mastering Chemistry</u> <u>Answers</u>

James N. Jensen

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