

HTML and CSS

6TH EDITION



INTRODUCTORY

Carey

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Patrick M. Carey



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positive m 1 table 16 3 three types of images formed by lenses the examples in figure 16 27 and figure 16 28 represent the three possible cases case 1 case 2 and case 3 summarized in table 16 3 in the table m is magnification the other symbols

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a a ray parallel to the principal axis of the lens is refracted through the principal focus b all rays of light refracted by the lens pass through the principal focus c the distance between the centre of the lens and the principal focus is the focal length d the principal focus of the lens is a point on the principal axis question 1 1

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web google classroom for the following redox equation $\text{MnO}_4^- + \text{H}_2\text{SO}_3 \rightarrow \text{Mn}^{2+} + \text{HSO}_4^-$ what will be the stoichiometric coefficient of H_2SO_3 if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN}_2^- + 2\text{e}^-$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- \rightarrow 2\text{AuCN}_2^- + 2\text{H}_2\text{O} + \text{O}_2$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add H^+ ions or the H_2O pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate $\text{Cr}_2\text{O}_7^{2-}$ and chloride Cl^- ions in acidic solution created by jay

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web balance each half reaction in basic solution $4\text{Fe}^{2+} + 5\text{H}_2\text{O} + \text{Cr}_2\text{O}_7^{2-} \rightarrow 4\text{Fe}^{3+} + 2\text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{TeO}_3^{2-} + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te} + 10\text{H}_2\text{O}$ $11\text{PbO}_2 + \text{IO}_3^- + 2\text{H}^+ \rightarrow \text{Pb}^{2+} + \text{IO}_3^- + \text{H}_2\text{SO}_4$

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ and OH^- in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 no 2 and $\text{C}_6\text{H}_8\text{O}_6$ b no 2 and $\text{C}_6\text{H}_8\text{O}_6$ h and no 2 c h and