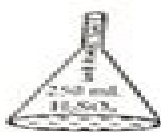


অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটাসিয়াম কপ্টার, এখানে মোট পটাসিয়াম, আরও কপ্টারের শীর্ষস্থানীয় কপ্টারের পটাসিয়াম এবং মনোমের টেম্পের পটাসিয়াম
পটাসিয়াম উত্তর অধ্যায়িকিত মেট্রিক হলে। এগুলোর অনুবর্তন করলে তুমি এ অধ্যায়িকিত থেকে পটাসিয়ামের পটাসিয়াম
এদের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেন্দ্র পটাসিয়াম লবণের প্রকণতা, 'অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করো।
- A-পাত্রের বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা পরিণত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলের প্রকণ বলে।

কেন্দ্র পটাসিয়াম লবণের প্রকণতা অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, আপমাত্রা বৃদ্ধির সাথে সাথে প্রকণতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট আপমাত্রার কেন্দ্র পটাসিয়াম লবণ নিষ্কৃত হয়। তখন প্রকণতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পটাসিয়ামে প্রকণিত করার সময় বিয়োজন প্রক্রিয়াটি অপমাত্রা বৃদ্ধি 32.4°C আপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি অপোৎপাদী হয়। এ কারণে 32.4°C আপমাত্রার উপরে আপমাত্রা বৃদ্ধিজনিত কারণে প্রকণতার হ্রাস ঘটে।

প্রশ্ন ২

আমরা জানি,
 $w = MV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
অতএব, প্রকণিত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
প্রকণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
প্রকণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$
প্রকণিত H_2SO_4 এর ভর, $w = ?$

প্রশ্ন-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রশ্ন-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

প্রশ্ন-A এর প্রকণের সাথে প্রশ্ন-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ প্রকণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{w}{v}$ HCl এর প্রকণ প্রকণ করে।

সি. কে. ১০/১০/১০

- নোড কাকে বলে?
- Cl_2 অপোনার— ব্যাখ্যা করো।
- উদাহরণের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদাহরণের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা এক হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন গঠিত হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোনার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপোনার তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোনার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহুণ ঘটে থাকে।

উদাহরণে বর্ণিত এলিটটি হলো 2% $\left(\frac{w}{v}\right)$ HCl।

2% $\left(\frac{w}{v}\right)$ HCl এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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Robert R. Redfield



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Table of Contents Hsc Chemistry 1st Paper Question 2solution

1. Understanding the eBook Hsc Chemistry 1st Paper Question 2solution
 - The Rise of Digital Reading Hsc Chemistry 1st Paper Question 2solution
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry 1st Paper Question 2solution
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry 1st Paper Question 2solution
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry 1st Paper Question 2solution
 - Personalized Recommendations

- Hsc Chemistry 1st Paper Question 2solution User Reviews and Ratings
- Hsc Chemistry 1st Paper Question 2solution and Bestseller Lists
- 5. Accessing Hsc Chemistry 1st Paper Question 2solution Free and Paid eBooks
 - Hsc Chemistry 1st Paper Question 2solution Public Domain eBooks
 - Hsc Chemistry 1st Paper Question 2solution eBook Subscription Services
 - Hsc Chemistry 1st Paper Question 2solution Budget-Friendly Options
- 6. Navigating Hsc Chemistry 1st Paper Question 2solution eBook Formats
 - ePub, PDF, MOBI, and More
 - Hsc Chemistry 1st Paper Question 2solution Compatibility with Devices
 - Hsc Chemistry 1st Paper Question 2solution Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Chemistry 1st Paper Question 2solution
 - Highlighting and Note-Taking Hsc Chemistry 1st Paper Question 2solution
 - Interactive Elements Hsc Chemistry 1st Paper Question 2solution
- 8. Staying Engaged with Hsc Chemistry 1st Paper Question 2solution
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Chemistry 1st Paper Question 2solution
- 9. Balancing eBooks and Physical Books Hsc Chemistry 1st Paper Question 2solution
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry 1st Paper Question 2solution
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Chemistry 1st Paper Question 2solution
 - Setting Reading Goals Hsc Chemistry 1st Paper Question 2solution
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Chemistry 1st Paper Question 2solution
 - Fact-Checking eBook Content of Hsc Chemistry 1st Paper Question 2solution

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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