UNIT III - THERMOCHEMISTRY

Calculate the H-F bond energy given the following data.

 $H-H(g) + F-F(g) \rightarrow 2 H-F(g)$ $\Delta H_{rxn} = -544 \text{ kJ}$

A. 1130 kJ/mol B. 565 kJ/mol C.)283 J/mol D. 300 kJ/mol E. 134 kJ/mol

1.0 g of the rocket fuel hydrazine (N2H4) is burned in a bomb calorimeter. The temperature
of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb
calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb
calorimeter.

A. 95.0 kJ/mol B. 2.96 kJ/mol C. 3.55 x 10³ B. 1.13 x 10⁵ kJ/mol

Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone O3.

 $NO(g) - O_3(g) \rightarrow NO_2(g) + O_2(g)$

Given the following data:

NO (g) + O (g) \rightarrow NO₂ (g) $\Delta H = -304.1 \text{ kJ}$ 2O₃ (g) \rightarrow 3O₂ (g) $\Delta H = -284.6 \text{ kJ}$ 2O (g) \rightarrow O₂ (g) $\Delta H = -495.0 \text{ kJ}$ A. -1083.7 kJ B. -693.8 kJ C. -198.8 kJ D -93.7 kJ E. -93.7 kJ

4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?

A. q<0, ΔE>0

D. q>0, ΔE<0

E. q>0, ΔE=0

C. q>0, ΔE>0

 M&M candies contain 209.0. Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?

A. 19.1 kJ B. 874 kJ S 1.91 x 10⁻² kJ D. 3.35 x 10⁵ kJ E. 335 kJ

Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.

2 C₄H₁₀ (l) + 13 O₂ (g) - 8 CO₂ (g) + 10 H₂O (l)

A. -90.7 kJ (B.)-98.3 kJ C. -49.2 kJ D. -108.4 kJ E. -196.6 kJ

7. When Karl Kaveman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.

A. 2.9 °C B. 14 °C C. 22 °C D. 25 °C E. 47 °C

General Chemistry Multiple Choice Questions And Answers

J Elliott

General Chemistry Multiple Choice Questions And Answers:

The book delves into General Chemistry Multiple Choice Questions And Answers. General Chemistry Multiple Choice Questions And Answers is a vital topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into General Chemistry Multiple Choice Questions And Answers, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 2: Essential Elements of General Chemistry Multiple Choice Questions And Answers
 - Chapter 3: General Chemistry Multiple Choice Questions And Answers in Everyday Life
 - Chapter 4: General Chemistry Multiple Choice Questions And Answers in Specific Contexts
 - \circ Chapter 5: Conclusion
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- 3. In chapter 2, the author will delve into the foundational concepts of General Chemistry Multiple Choice Questions And Answers. The second chapter will elucidate the essential principles that need to be understood to grasp General Chemistry Multiple Choice Questions And Answers in its entirety.
- 4. In chapter 3, this book will examine the practical applications of General Chemistry Multiple Choice Questions And Answers in daily life. This chapter will showcase real-world examples of how General Chemistry Multiple Choice Questions And Answers can be effectively utilized in everyday scenarios.
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- 6. In chapter 5, this book will draw a conclusion about General Chemistry Multiple Choice Questions And Answers. This chapter will summarize the key points that have been discussed throughout the book.
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