



Figure 3. 7E03 gas turbine borescope inspection access locations

Boreprobe inspection access locations for F-class gas turbines can be found in Appendix E.

Figure 4 provides a recommended interval for a planned boreprobe inspection program following initial baseline inspections. It should be recognized that these boreprobe inspection intervals are based on average unit operating modes. Adjustment of these boreprobe intervals may be made based on operating experience, mode of operation, fuels used, and the results of previous boreprobe inspections.

In general, an annual or semiannual boreprobe inspection uses all the available access points to verify the condition of the internal hardware. This should include, but is not limited to, signs of excessive gas path fouling, symptoms of surface degradation (such as erosion, corrosion, or spalling), displaced components, deformation or object damage, material loss, nicks, dents, cracking, indications of contact or rubbing, or other anomalous conditions.

Boreprobe	Gas and Distillate Fuel Oil	At combustion inspection or annually, whichever occurs first
	Heavy Fuel Oil	At combustion inspection or semiannually, whichever occurs first

Figure 4. Boreprobe inspection planning

During BIs and similar inspections, the condition of the upstream components should be verified, including all systems from the filter house to the compressor inlet.

The application of a boreprobe monitoring program will assist with the scheduling of outages and preplanning of parts requirements, resulting in outage preparedness, lower maintenance costs, and higher availability and reliability of the gas turbine.

Ge Frame 5 Service Manual

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web identifying chemical reaction types key part 1 identify the type of reaction 1 $\text{Na}_3\text{PO}_4 + 3\text{KOH} \rightarrow 3\text{NaOH} + \text{K}_3\text{PO}_4$ double displacement 2 $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2\text{LiCl}$ double displacement 3 $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$ combustion 4 $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ single displacement 5 $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ decomposition 6 $\text{P}_4 + 3\text{O}_2 \rightarrow 2\text{P}_2\text{O}_3$

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web chemical equations can also be used to represent physical processes write a chemical reaction for the boiling of water including the proper phase labels chemical equations can also be used to represent physical processes write a chemical reaction for the freezing of water including the proper phase labels explain why $4\text{Na} + \text{S} \rightarrow 2\text{Na}_2\text{S}$

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web name score date balancing and identifying chemical reactions a identify each type of chemical reactions 1 $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ 2 $\text{Sr} + 2\text{AgNO}_3 \rightarrow \text{Sr}(\text{NO}_3)_2 + 2\text{Ag}$ 3 $2\text{LiBr} + \text{F}_2 \rightarrow 2\text{LiF} + \text{Br}_2$ 4 $2\text{Na} + \text{F}_2 \rightarrow 2\text{NaF}$ 5 $2\text{Br}_2 + \text{H}_2 \rightarrow 2\text{HBr}$ b balance the chemical reactions and identify each type

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web 1 $3\text{NaBr} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3\text{HBr}$ type of reaction double displacement 2 $3\text{Ca}(\text{OH})_2 + 2\text{Al}_2(\text{SO}_4)_3 \rightarrow 3\text{CaSO}_4 + 4\text{Al}(\text{OH})_3$ type of reaction double displacement 3 $3\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow 2\text{Fe} + 3\text{MgO}$ type of reaction single displacement 4 $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ type of reaction combustion 5 $2\text{PbSO}_4 + 2\text{PbSO}_3 + \text{O}_2 \rightarrow$ type of reaction

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web a chemical reaction is a process in which one or more substances are converted to one or more different substances the starting substances are called the reactants and the new substances that form are called the products

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web study guide chemical reactions 1 give an example of a chemical reaction leaves turning color fireworks food spoiling 2 differentiate between a physical and chemical change physical does not change the identity of the matter can be changed back easily chemical new substance is created with new physical and chemical