

Experiment 10

Oxidation-Reduction (Redox) Reactions

Pre-Lab Assignment

Before coming to lab:

- Read the lab thoroughly.
- Answer the pre-lab questions that appear at the end of this lab exercise.

Purpose

A set of oxidation-reduction (redox) reactions will be performed to determine the relative strengths as reducing and oxidizing agents of metals and other ions. An activation series will be devised and used to predict whether a reaction will occur spontaneously or not.

Background

Oxidation-reduction (redox) reactions are a classification of chemical changes that involve the transfer of electrons. An example of a redox reaction is shown in Eqn. 1, when magnesium metal reacts with chlorine gas.



At first glance, the electron exchange is not apparent as electrons are not included in the overall balanced equation. However, electrons are shown when the reaction is split into halves: an oxidized half-reaction and a reduced half-reaction, as in Eqn. 2 and Eqn. 3 respectively.



Notice that in the half-reactions, the electrons are required to balance the charges. Since Mg is becoming more positive, one magnesium atom loses two electrons to become a cation and is said to be **oxidized**. Chlorine is becoming more negative, so one molecule of $\text{Cl}_2(\text{g})$ gains two electrons to become two chloride anions and is said to be **reduced**. Both oxidation and reduction happen simultaneously in a single redox reaction. Overall, Mg is transferring two electrons to chlorine (Fig. 1).

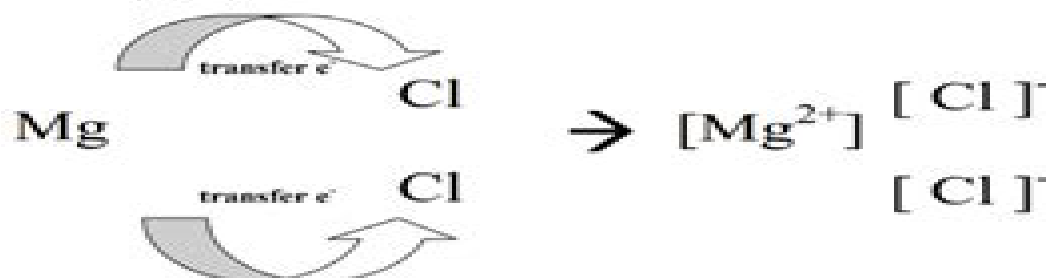


Fig. 1: Redox reaction between Mg(s) and Cl₂(g) to produce MgCl₂(aq)

Experiment 46 Oxidation Reduction Reactions Answers

JS Bruner



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AP Chemistry For Dummies Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam Gearing up for the AP Chemistry exam AP Chemistry For Dummies is packed with all the resources and help you need to do your very best Focused on the chemistry concepts and problems the College Board wants you to know this AP Chemistry study guide gives you winning test taking tips multiple choice strategies and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day This user friendly guide helps you prepare without perspiration by developing a pre test plan organizing your study time and getting the most out of your AP course You ll get help understanding atomic structure and bonding grasping atomic geometry understanding how colliding particles produce states and so much more To provide students with hands on experience AP chemistry courses include extensive labwork as part of the standard curriculum This is why the book dedicates a chapter to providing a brief review of common laboratory equipment and techniques and another to a complete survey of recommended AP chemistry experiments Two full length practice exams help you build your confidence get comfortable with test formats identify your strengths and weaknesses and focus your studies You ll discover how to Create and follow a pretest plan Understand everything you must know about the exam Develop a multiple choice strategy Figure out displacement combustion and acid base reactions Get familiar with stoichiometry Describe patterns and predict properties Get a handle on organic chemistry nomenclature Know your way around laboratory concepts tasks equipment and safety Analyze laboratory data Use practice exams to maximize your score Additionally you ll have a chance to brush up on the math skills that will help you on the exam learn the critical types of chemistry problems and become familiar with the annoying exceptions to chemistry rules Get your own copy of AP Chemistry For Dummies to build your confidence and test taking know how so you can ace that exam

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appendix listing science supply companies a glossary further reading with books and websites and an index

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Electrochemical Energy Pei Kang Shen, Chao-Yang Wang, San Ping Jiang, Xueliang Sun, Jiujun Zhang, 2018-10-08 Electrochemical Energy Advanced Materials and Technologies

covers the development of advanced materials and technologies for electrochemical energy conversion and storage The book was created by participants of the International Conference on Electrochemical Materials and Technologies for Clean Sustainable Energy ICES 2013 held in Guangzhou China and incorporates select papers presented at the conference More than 300 attendees from across the globe participated in ICES 2013 and gave presentations in six major themes Fuel cells and hydrogen energy Lithium batteries and advanced secondary batteries Green energy for a clean environment Photo Electrocatalysis Supercapacitors Electrochemical clean energy applications and markets Comprised of eight sections this book includes 25 chapters featuring highlights from the conference and covering every facet of synthesis characterization and performance evaluation of the advanced materials for electrochemical energy It thoroughly describes electrochemical energy conversion and storage technologies such as batteries fuel cells supercapacitors hydrogen generation and their associated materials The book contains a number of topics that include electrochemical processes materials components assembly and manufacturing and degradation mechanisms It also addresses challenges related to cost and performance provides varying perspectives and emphasizes existing and emerging solutions The result of a conference encouraging enhanced research collaboration among members of the electrochemical energy community Electrochemical Energy Advanced Materials and Technologies is dedicated to the development of advanced materials and technologies for electrochemical energy conversion and storage and details the technologies current achievements and future directions in the field

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