4. a. A sample of a CaCl₂2H₂O/K₂C₂O₂H₂O solid salt mixture is dissolved in -150 mL of desonated water previously adjusted to a pH that is basic. The precipitate, after having been filtered, was sur-dried and weighed. Data for Trial I were obtained as shown. Complete the following table. (See Report Short.) Record calculated values with the context number of significant figures.

Colombotion Zome A. Precipitation of CaC,O,H,O from the Salt Misture Data Analysis, L. 0.879 3. Mass of salt mixture (a) 1.896 4. Mass of filter paper (g) Mass of filter paper and CaC₂O₂·H₂O (g). 2.180 6. Mass of air-dried CaC-O, H-O (g) B. Determination of Limiting Reactant Data Amalysis, 3. CarCiti-2H,O 1. Limiting reactant in salt mixture 2. Excess reactant in salt mixture Duta Analysis 1. Moles of CaC₂O₂H₂O precipitated (mol) Data Analysis, 4. Ohne male allegations 2. Moles of limiting reactant in salt mixture (mol) See expandion 8.1. 3. Mass of limiting reactant in salt mixture (g) Obcom controllation. 4. Mass of excess reactant in salt mixture (g) Data Analysis, 5. Ecoulty mans of east mixture minus mass of limiting maction. Ships endoutation. 5. Percent limiting reactant in salt mixture (%) Show daleylation. 6. Percent excess reactant in salt mixture (%)

4. b. For Trials 2 and 3, the percent CaClg2H₂O in the mixture was 39.7% and 28.4% respectively.

a. What is the average percent of CaCl, 2H,O in the sample? Data Analysis, B.

b. What are the standard deviation and the relative standard deviation (%RSD) for the percent of CaCl₂·2H₂O in the sample? Data Analysis, C and D.

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