

$$\Delta S [\text{H}_2\text{O}_2 (g)] = 232.9 \text{ J K}^{-1} \text{ mol}^{-1} \quad \Delta H_f [\text{H}_2\text{O}_2 (g)] = -136.1 \text{ kJ mol}^{-1}$$

$$\Delta S_{\text{universe}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}} = 5.3 \text{ J K}^{-1} \text{ mol}^{-1} \text{ spontaneous}$$

7. The reaction $\text{NH}_3(g) + \text{HCl}(g) \rightarrow \text{NH}_4\text{Cl}(s)$ is spontaneous at room temperature. Determine the temperature at which it is no longer spontaneous. Given information

$$S^\circ [\text{NH}_3(g)] = 193.0 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$S^\circ [\text{HCl}(g)] = 187.0 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$S^\circ [\text{NH}_4\text{Cl}(s)] = 94.56 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$\Delta H_f^\circ [\text{NH}_3(g)] = -46.3 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\circ [\text{HCl}(g)] = -92.3 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\circ [\text{NH}_4\text{Cl}(s)] = -315.39 \text{ kJ mol}^{-1}$$

$$T < 619 \text{ K}$$

8. Calculate the change in the entropies of the system and the surroundings, and the total change in entropy, when the volume of a sample of argon gas of mass 21 g at 298 K and 1.50 bar increases from 1.20 dm³ to 4.60 dm³ in (a) an isothermal reversible expansion, (b) an isothermal irreversible expansion against $p_{\text{ex}} = 0$, and (c) an adiabatic reversible expansion

$$\text{a) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = 0 \text{ J/K}$$

$$\text{b) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = +3.0 \text{ J/K}$$

$$\text{c) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = +3.0 \text{ J/K}$$

9. The standard molar enthalpy of formation of $\text{H}_2\text{O}(g)$ is $-241.8 \text{ kJ mol}^{-1}$ and at 298 K, water has

entropy S° given by the following equation (see also the coefficients given in Table 2.1). Calculate the standard molar entropy of $\text{H}_2\text{O}(g)$ at 298 K and 1 bar.

$$\text{a) } 241.8 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$\text{b) } 241.8 \text{ J K}^{-1} \text{ mol}^{-1}$$

10. The standard molar heat capacity $C_{p,m}$ of a diatomic gas at room temperature. The molar enthalpy of formation of $\text{H}_2\text{O}(g)$ is $-241.8 \text{ kJ mol}^{-1}$ and the molar enthalpy of formation of $\text{H}_2\text{O}(l)$ is $-285.8 \text{ kJ mol}^{-1}$. What is the standard change in the system entropy ΔS° of a diatomic gas at the room temperature?

$$\text{a) } 1.24 \text{ J K}^{-1} \text{ mol}^{-1}$$

11. Suppose water and the gas mixture is more negative or more negative than 0. The standard molar enthalpy of formation of $\text{H}_2\text{O}(g)$ is $-241.8 \text{ kJ mol}^{-1}$.

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C. A. Trapp, M. P. Cady, Carmen Giunta



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Table of Contents Exercises For Thermodynamics Answers

1. Understanding the eBook Exercises For Thermodynamics Answers
 - The Rise of Digital Reading Exercises For Thermodynamics Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Exercises For Thermodynamics Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Exercises For Thermodynamics Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Exercises For Thermodynamics Answers
 - Personalized Recommendations
 - Exercises For Thermodynamics Answers User Reviews and Ratings
 - Exercises For Thermodynamics Answers and Bestseller Lists
5. Accessing Exercises For Thermodynamics Answers Free and Paid eBooks
 - Exercises For Thermodynamics Answers Public Domain eBooks
 - Exercises For Thermodynamics Answers eBook Subscription Services

- Exercises For Thermodynamics Answers Budget-Friendly Options
- 6. Navigating Exercises For Thermodynamics Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Exercises For Thermodynamics Answers Compatibility with Devices
 - Exercises For Thermodynamics Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Exercises For Thermodynamics Answers
 - Highlighting and Note-Taking Exercises For Thermodynamics Answers
 - Interactive Elements Exercises For Thermodynamics Answers
- 8. Staying Engaged with Exercises For Thermodynamics Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Exercises For Thermodynamics Answers
- 9. Balancing eBooks and Physical Books Exercises For Thermodynamics Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Exercises For Thermodynamics Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Exercises For Thermodynamics Answers
 - Setting Reading Goals Exercises For Thermodynamics Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Exercises For Thermodynamics Answers
 - Fact-Checking eBook Content of Exercises For Thermodynamics Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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