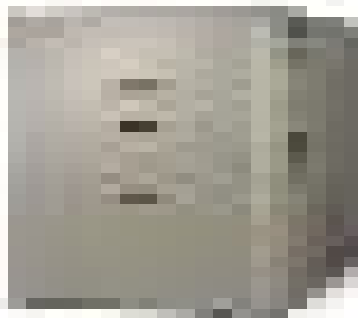


LABORATORY

EXPERIMENT 1
PURPOSE
OBJECTIVES
THEORY
PROCEDURE



RESULTS
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INTRODUCTION
The purpose of this experiment is to determine the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction is catalyzed by the presence of a small amount of potassium iodide. The rate of reaction is measured by the volume of oxygen gas evolved over a period of time. The experiment is carried out at three different temperatures: 25°C, 35°C, and 45°C. The results show that the rate of reaction increases with increasing temperature. This is due to the fact that at higher temperatures, the molecules have more kinetic energy and are therefore more likely to collide with sufficient energy to overcome the activation energy barrier. The experiment also demonstrates the effect of the concentration of the reactants on the rate of reaction. It is found that the rate of reaction increases with increasing concentration of both hydrogen peroxide and potassium iodide. This is because a higher concentration of reactants leads to a higher frequency of collisions between the molecules. The experiment is carried out in a series of test tubes, and the volume of oxygen gas evolved is measured using a gas syringe. The results are recorded in a table, and a graph is plotted showing the rate of reaction versus temperature. The graph shows a clear upward trend, indicating that the rate of reaction increases with increasing temperature. The experiment is repeated three times to ensure the accuracy of the results. The average values are then used to draw conclusions about the effect of temperature on the rate of reaction. The experiment is a classic demonstration of the effect of temperature on the rate of reaction, and it provides a good opportunity for students to learn about the factors that affect the rate of a chemical reaction. The experiment is also a good example of how to carry out a scientific investigation, from the formulation of a hypothesis to the collection of data and the drawing of conclusions. The experiment is suitable for students at the high school or college level, and it can be adapted to suit the needs of different classes. The experiment is a valuable learning experience for students, and it helps them to develop their understanding of the principles of chemistry. The experiment is also a good example of how to use laboratory equipment, and it helps students to learn about the safety procedures that should be followed when working with chemicals. The experiment is a great way to engage students in learning, and it helps them to develop their critical thinking skills. The experiment is a valuable addition to any chemistry curriculum, and it is a great way to teach students about the factors that affect the rate of a chemical reaction. The experiment is a classic demonstration of the effect of temperature on the rate of reaction, and it provides a good opportunity for students to learn about the factors that affect the rate of a chemical reaction. The experiment is also a good example of how to carry out a scientific investigation, from the formulation of a hypothesis to the collection of data and the drawing of conclusions. The experiment is suitable for students at the high school or college level, and it can be adapted to suit the needs of different classes. The experiment is a valuable learning experience for students, and it helps them to develop their understanding of the principles of chemistry. The experiment is also a good example of how to use laboratory equipment, and it helps students to learn about the safety procedures that should be followed when working with chemicals. The experiment is a great way to engage students in learning, and it helps them to develop their critical thinking skills. The experiment is a valuable addition to any chemistry curriculum, and it is a great way to teach students about the factors that affect the rate of a chemical reaction.

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L Darling-Hammond



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