

11 Which compound has the greatest percent composition by mass of sulfur?

- (1) BaS (2) CaS (3) MgS (4) SrS

way #64

12 Two molecules of HBr collide and then form  $H_2$  and  $Br_2$ . During the collision, the bonds in the HBr molecules are

- (1) broken as energy is absorbed  
(2) broken as energy is released  
(3) formed as energy is absorbed  
(4) formed as energy is released

way #43

BA/RF

13 Which atom in the ground state has a stable electron configuration?

- (1) carbon (2) magnesium (3) neon (4) oxygen

way #49

14 Which statement describes a multiple covalent bond?

- (1) Two electrons are shared.  
(2) Four electrons are shared.  
(3) Two electrons are transferred.  
(4) Four electrons are transferred.

way #46

15 The electronegativity difference between the atoms in a molecule of HCl can be used to determine

- (1) the entropy of the atoms  
(2) the atomic number of the atoms  
(3) the first ionization energy of the atoms  
(4) the polarity of the bond between the two atoms

way #51

16 Which two gases can not be broken down by chemical means?

- (1) CO and He (2) CO and  $NH_3$  (3) Xe and He (4) Xe and  $NH_3$

way #5

elements Table S

17 Two substances in a mixture differ in density and particle size. These properties can be used to

- (1) separate the substances  
(2) chemically combine the substances  
(3) determine the freezing point of the mixture  
(4) predict the electrical conductivity of the mixture

18 Which unit is used to express an amount of thermal energy?

- (1) gram (2) mole (3) joule (4) pascal

Table D

19 Under which conditions of temperature and pressure does a real gas behave most like an ideal gas?

- (1) low temperature and low pressure  
(2) low temperature and high pressure  
(3) high temperature and low pressure  
(4) high temperature and high pressure

way #9

20 According to the kinetic molecular theory for an ideal gas, all gas particles

- (1) are in random, constant, straight-line motion  
(2) are separated by very small distances relative to their sizes  
(3) have strong intermolecular forces  
(4) have collisions that decrease the total energy of the system

way #8

21 Which mathematical expression represents the heat of reaction for a chemical reaction?

- (1) (the heat of fusion) - (the heat of vaporization)  
(2) (the heat of vaporization) - (the heat of fusion)  
(3) (the potential energy of the products) - (the potential energy of the reactants)  
(4) (the potential energy of the reactants) - (the potential energy of the products)

way #78

22 At 101.3 kPa and 298 K, a 1.0-mole sample of which compound absorbs the greatest amount of heat as the entire sample dissolves in water?

- (1) LiBr -48.83 (2) NaCl +2.38 (3) NaOH -46.51 (4)  $NH_4Cl$  +14.8

way #11

Table I

23 For a reaction at equilibrium, which change can increase the rates of the forward and reverse reactions?

- (1) a decrease in the concentration of the reactants  
(2) a decrease in the surface area of the products  
(3) an increase in the temperature of the system  
(4) an increase in the activation energy of the forward reaction

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**Brendan G. Carr**



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## **Table of Contents June 2008 Physics Regents Answers**

1. Understanding the eBook June 2008 Physics Regents Answers
  - The Rise of Digital Reading June 2008 Physics Regents Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying June 2008 Physics Regents Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an June 2008 Physics Regents Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from June 2008 Physics Regents Answers
  - Personalized Recommendations
  - June 2008 Physics Regents Answers User Reviews and Ratings
  - June 2008 Physics Regents Answers and Bestseller Lists
5. Accessing June 2008 Physics Regents Answers Free and Paid eBooks
  - June 2008 Physics Regents Answers Public Domain eBooks
  - June 2008 Physics Regents Answers eBook Subscription Services
  - June 2008 Physics Regents Answers Budget-Friendly Options

6. Navigating June 2008 Physics Regents Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - June 2008 Physics Regents Answers Compatibility with Devices
  - June 2008 Physics Regents Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of June 2008 Physics Regents Answers
  - Highlighting and Note-Taking June 2008 Physics Regents Answers
  - Interactive Elements June 2008 Physics Regents Answers
8. Staying Engaged with June 2008 Physics Regents Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers June 2008 Physics Regents Answers
9. Balancing eBooks and Physical Books June 2008 Physics Regents Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection June 2008 Physics Regents Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine June 2008 Physics Regents Answers
  - Setting Reading Goals June 2008 Physics Regents Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of June 2008 Physics Regents Answers
  - Fact-Checking eBook Content of June 2008 Physics Regents Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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