



MID-TERM EXAMINATIONS

Class : X (CBSE)

Sub : General Science

2023-24

Max. Marks : 80

Time : 3 Hrs.

General Instructions:

- This question paper consists of 39 questions in 5 sections.
- All questions are compulsory. However, an internal choice is provided in some questions. A student is expected to attempt only one of these questions.
- Section A consists of 20 objective type questions carrying 1 mark each.
- Section B consists of 6 Very Short questions carrying 02 marks each. Answers to these questions should be in the range of 30 to 50 words.
- Section C consists of 7 Short Answer type questions carrying 03 marks each. Answers to these questions should be in the range of 50 to 80 words.
- Section D consists of 3 Long Answer type questions carrying 05 marks each. Answer to these questions should be in the range of 80 to 120 words.
- Section E consists of 3 source-based/case-based units of assessment of 04 marks each with sub-parts.

SECTION - A

Select and write the most appropriate option out of the four options given for each of the questions 1 - 20. There is no negative mark for incorrect response. $20 \times 1 = 20$

- Three beakers labelled as A, B and C each containing 25 ml. of water were taken. A small amount of NaOH , anhydrous CuSO_4 and NH_4Cl were added to the beakers A, B and C respectively. It was observed that there was an increase in the temperature of the solutions contained in beakers A and B, whereas in case of beaker C, the temperature of the solution falls. Which one of the following statement (s) is (are) correct []
 - In beakers A and B exothermic process
 - In beakers A and B, endothermic process has occurred
 - In beaker C exothermic process has occurred
 - In beaker C endothermic process has occurred
 a) (i) only b) (ii) only c) (i) and (iv) d) (ii) and (iii)
- Which of the following reaction is characterised by the yellow colour of product? []
 - $\text{Zn}(s) + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$
 - $\text{Na}_2\text{CO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
 - $2\text{PbO}_2 \xrightarrow{\Delta} 2\text{PbO}(s) + \text{O}_2(g)$
 - $2\text{KClO}_3 \xrightarrow{\Delta} 2\text{KCl} + 3\text{O}_2(g)$
- Which of the following statements about the given reaction are correct? []

$$3\text{Fe}_{(s)} + 4\text{H}_2\text{O}_{(g)} \rightarrow \text{Fe}_3\text{O}_{4(s)} + 4\text{H}_{2(g)}$$
 - Iron metal is getting oxidised
 - Water is getting reduced
 - Water is acting as reducing agent
 - Water is acting as oxidising agent
 a) (i), (ii) and (iii) b) (iii) and (iv) c) (i) and (iv) d) (ii) and (iv)

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