Name of the student

Date – 14.09.2022 (Wednesday) Section

GENERAL INSTRUCTIONS:

- All questions are compulsory.
- Question nos. 1 to 21 are MCQs with one correct option and carrying 1 mark each.
- Question nos. 22 to 24 are objective/short answer type questions.
- Question nos. 25 to 28 are of 2 marks each.
- Question nos. 29 to 32 is of 3 marks each.
- Question nos. 33 to 35 is of 5 marks each.

SECTION - A

Q. 1 to 21 is of 1 mark each.

- Q.1 Which of the following aqueous solution should have highest boiling point?
 - a) 1 M NaOH
- b) 1 M Na₂SO₄
- c) 1M NH₄NO₃
- d) 1 M KNO₃

- Q.2 Which of the following statement is false?
 - a) Two different solution of sucrose of same molality prepare in different solvent will have the same depression in freezing point.
 - b) The osmotic pressure of a solution is given by the equation π =CRT. (where C is the molarity of the solution)
 - c) Decreasing order of osmotic pressure for 0.01M aqueous solution of Barium Chloride, Potassium chloride, acetic acid and sucrose is BaCl₂ > KCl > CH₃COOH > Sucrose
 - d) According to Raoult's Law, the vapour pressure exeterd by a volatile component of solution is directly proportional to its molar fraction in the solution.
- Q.3 We have three aqueous solution of NaCl labelled as 'A', 'B' and 'C' with concentration 0.1 M, 0.01 M and 0.001 M respectively. The value of Van't Hoff factor for these solution will be in the order
 - a) ia < in < ic
- a) $i_A > i_B > i_C$
- a) $i_A = i_B = i_C$

- a) iA < iB > ic
- Q.4 On the basis of the information given below mark the correct options -
 - In bromoethane and chloroethane mixture intermolecular interaction of A-A and B-B type are nearly same as A-B type interaction.
 - ii) In the ethanol and acetone mixture A-A or B-B type intermolecular interactions are stronger than A-B type interaction.

Hsc Half Yearly Past Papers And Chemistry

Michael Hoelscher

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