

## Basic Stoichiometry

## Why?

In this activity we will address the question:

*How do I convert between moles of different chemical species in a given reaction?*

## Model 1



2 mol A produces 5 mol C

3 mol B produces 4 mol D

2 mol A reacts with 3 mol B

4 mol A produces 10 mol C

6 mol B produces 8 mol D

6 mol A reacts with 9 mol B

1. What do the coefficients of a balanced equation represent?

3. If 15 mol B are reacted, how many mol C are formed? Show your work.

2. If 6 mol A are reacted, how many mol D are formed? Explain IN WORDS how you figured this out.

4. If 8 mol of A are reacted, how many mol B are used? Show your work.

## Model 2



If 3 mol A react, how many mol C are produced?

$$3 \text{ mol A} \times \frac{5 \text{ mol C}}{2 \text{ mol A}} = 7.5 \text{ mol C}$$

$$\text{In general: } \text{mol known} \times \frac{\text{mol unknown}}{\text{mol known}} = \text{mol unknown}$$

5. If 21.0 mol A are reacted, how many mol C are produced? Using dimensional analysis, show how you calculated your answer.

6. If 21.0 mol B are reacted, how many mol A must have reacted?

# Hs Chemistry Pogil Activity Basic Stoichiometry Answers

**Dharshi Bopegedera**



## **Hs Chemistry Pogil Activity Basic Stoichiometry Answers:**

## The Enigmatic Realm of **Hs Chemistry Pogil Activity Basic Stoichiometry Answers**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing in short supply of extraordinary. Within the captivating pages of **Hs Chemistry Pogil Activity Basic Stoichiometry Answers** a literary masterpiece penned with a renowned author, readers embark on a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting impact on the hearts and minds of people who partake in its reading experience.

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3\text{ mole H}_2\text{S} \times \frac{2\text{ mole H}_2\text{O}}{2\text{ mole H}_2\text{S}} \times 18\text{g H}_2\text{O} = 149.4\text{g H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained / calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$  c determine the percent yield yield actual  $\times 100$   $260.0 \text{ g} \times 100 / 86.01$  theoretical 302.3 g

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield / actual yield  $\times 100$  based on this definition we would expect a percent yield

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