

Questions and Problems

Q1 According to your graph, what is the relationship between the pressure and the volume of a gas?

* As the volume of the gas increases the pressure of the gas decreases. They have inverse proportional relationship.

Q2 On your graph, what is the volume of the gas at a pressure of 760. mmHg?

* After plotting the data to excel, the graph produced a linear equation $y = -28.873x + 1539.3$ where y is the pressure and x is the volume of gas.

We need to find x at 760 mmHg.

Hence,

$$760 \text{ mmHg} = -28.873x + 1539.8 \Rightarrow x = 26.99 \text{ mL}$$

Therefore, at 760 mmHg the volume of the gas is 26.99 mL

Q3 On your graph, what is the pressure of the gas when the volume is 30.0 mL?

* $x = 30.0 \text{ mL}$, $y = ?$

$$y = -28.873 (30.0) + 1539.3$$

$$y = 673.1 \text{ mmHg}$$

Therefore, when the volume of the gas is 30.0 mL the pressure is 673.1 mmHg

Q4 A sample of helium has a volume of 325 mL and a pressure of 655 mmHg. What will be the pressure, in mmHg, if the sample of helium is compressed to 125 mL. (T, n constant)?

$$P_1 V_1 = P_2 V_2$$

$$P_2 = (P_1 V_1) \div V_2 = (655 \text{ mmHg} \times 325 \text{ mL}) \div 125 \text{ mL} = 1,703 \text{ mmHg}$$

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Table of Contents Holt Chemistry Section Review Answers Gases

1. Understanding the eBook Holt Chemistry Section Review Answers Gases
 - The Rise of Digital Reading Holt Chemistry Section Review Answers Gases
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Section Review Answers Gases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Section Review Answers Gases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Section Review Answers Gases
 - Personalized Recommendations
 - Holt Chemistry Section Review Answers Gases User Reviews and Ratings
 - Holt Chemistry Section Review Answers Gases and Bestseller Lists
5. Accessing Holt Chemistry Section Review Answers Gases Free and Paid eBooks
 - Holt Chemistry Section Review Answers Gases Public Domain eBooks
 - Holt Chemistry Section Review Answers Gases eBook Subscription Services
 - Holt Chemistry Section Review Answers Gases Budget-Friendly Options
6. Navigating Holt Chemistry Section Review Answers Gases eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Chemistry Section Review Answers Gases Compatibility with Devices
 - Holt Chemistry Section Review Answers Gases Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Section Review Answers Gases
 - Highlighting and Note-Taking Holt Chemistry Section Review Answers Gases
 - Interactive Elements Holt Chemistry Section Review Answers Gases
8. Staying Engaged with Holt Chemistry Section Review Answers Gases

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Holt Chemistry Section Review Answers Gases
- 9. Balancing eBooks and Physical Books Holt Chemistry Section Review Answers Gases
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Section Review Answers Gases
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Holt Chemistry Section Review Answers Gases
 - Setting Reading Goals Holt Chemistry Section Review Answers Gases
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Holt Chemistry Section Review Answers Gases
 - Fact-Checking eBook Content of Holt Chemistry Section Review Answers Gases
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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