CHAPTER 19 REVIEW

Oxidation-Reduction Reactions

SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

- All the following equations involve redex reactions except
 - (a) CuO + H₂O → Ca(OH)₂.
 - (b) 2SO₂ + O₂ → 2SO₃.
 - (c) $2HgO \rightarrow 2Hg + O_2$.
 - (d) SnCl₄ + 2FeCl₂ → 2FeCl₃ + SnCl₂.
- Assign the correct exidation number to the individual atom or ion below.
 - _____ a. Ma in MaO₂
 - _____ b. S in S_R
 - e. Cl in CaCl₂
 - ______ d. 1 is 10₃
 - ______ e. C in HCO₃
 - f. Fe in Fe₂(SO₄)₃
 - _____ g. S in Fe₂(SO₄)₃
- 3. In each of the following half-reactions, determine the value of x.
 - _____ a. $S^{6+} + x e^- \rightarrow S^{2-}$
 - _____ b. $2Br'' \rightarrow Br_2 + 2e^-$
 - _____ c. $Sn^{4+} + 2e^- \rightarrow Sn^a$
 - d. Which of the above half-reactions represent reduction processes?
- 4. Give examples, other than those listed in Table 1 on page 631 of the text, for the following:
 - ______ n. a compound containing H in a −1 oxidation state
 - ______ b. a peroxide
 - e. a polyatomic son in which the oxidation number for S is +4
 - d. a substance in which the excidation number for F is not −1

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K Morrison

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