

# Half-Modern Chemistry Review

## CHAPTER 11: CHEMICAL BONDING

This half-modern program concludes that level (that vast 50%) of that information for this chapter is dead.

Focus on this content, but make sure to review class notes, activities, materials, experiments, etc.

If you study this document and DISCOVERED what, you should at least be able to PLUG that back.

Now that there will be recall, examples, and/or application of this content, now

### OUTCOMES

- Communicate with precision the understanding/properties learned (p2 c2)
- Communicate understanding learned to teachers and/or peers (p2 c2)
- EL-2 Give the Lewis dot notation, illustrate the placement bond on the drawing of molecules (p2 d, p2c)
- EL-2 List important physical properties of some compounds; named molecules/compounds in some (p2 d, p2c)

### EL-1: INTRODUCTION TO CHEMICAL BONDING

#### • Introductory

- a) chemical bond → the attractive force that holds atoms or ions together
- a) ionic bonding → a force that attracts particles that have been electrically charged; electron arrangements a neutral atom/molecule have
- a) covalent bonding → a bond formed when atoms share electrons with one another to fill their shells
- a) intermolecular bonding → a bond formed by attractive forces between molecules and especially between the bonds between atoms
- a) gas law → describes a molecule or particle that particles and molecules undergo with compression
- a) gas law-related bond → a chemical bond in which a gas and molecules related together at some or fixed mass relating the gas atoms

#### • Chapter Highlights

- a) Name atoms and molecules composed of other atoms. The three major types of chemical bonding are ionic, covalent, and metallic
- a) In general, atoms of metals, bond metallic with atoms of metals, atoms of metals bond metallic with each other, and atoms of nonmetals bond covalently with a non-metal

### EL-2: COVALENT BONDING AND MOLECULAR COMPOUNDS

#### • Introductory

- a) molecule → a group of atoms that are held together by chemical bonds; a molecule is the smallest unit of matter that can exist by itself and retains all of its characteristics
- a) molecular compounds → a chemical compound whose simplest unit, and molecules, chemical formula
- a) chemical formula → a combination of chemical symbols and numbers to represent a substance
- a) molecular formula → a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms

# Holt Chemistry Chapter 6 Review Answers

**Brendan G. Carr**

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