

Chapter 10: Equilibrium

10.1. What is the equilibrium constant for the reaction of H_2O and CO_2 to form H_2CO_3 ?

$$\text{H}_2\text{O} + \text{CO}_2 \rightleftharpoons \text{H}_2\text{CO}_3 \quad K_c = \frac{[\text{H}_2\text{CO}_3]}{[\text{H}_2\text{O}][\text{CO}_2]} = 1.5 \times 10^{-3}$$

- H_2O
- CO_2
- H_2CO_3

10.2. The equilibrium constant for the reaction of H_2O and CO_2 to form H_2CO_3 is 1.5×10^{-3} . What is the concentration of H_2CO_3 if the concentration of H_2O is 55.5 M and the concentration of CO_2 is 0.033 M ?

- 1.5×10^{-3}
- 55.5 M
- 0.033 M
- $1.5 \times 10^{-3} \text{ M}$

10.3. The equilibrium constant for the reaction of H_2O and CO_2 to form H_2CO_3 is 1.5×10^{-3} . What is the concentration of H_2CO_3 if the concentration of H_2O is 55.5 M and the concentration of CO_2 is 0.033 M ?

- 1.5×10^{-3}
- 55.5 M
- 0.033 M
- $1.5 \times 10^{-3} \text{ M}$

$$K_c = \frac{[\text{H}_2\text{CO}_3]}{[\text{H}_2\text{O}][\text{CO}_2]} \Rightarrow [\text{H}_2\text{CO}_3] = K_c [\text{H}_2\text{O}][\text{CO}_2] = 1.5 \times 10^{-3} (55.5)(0.033) = 2.7 \times 10^{-3} \text{ M}$$

10.4. Which of the following is a property of a system at equilibrium?

- The rate of the forward reaction is equal to the rate of the reverse reaction.
- The concentration of the reactants is equal to the concentration of the products.
- The system is at a minimum of energy.
- The system is at a maximum of energy.

10.5. A reaction is exothermic. If the temperature is increased, the equilibrium constant K_c will:

- increase.
- decrease.
- remain the same.
- be zero.

$$\text{Exothermic: } K_c \text{ decreases as } T \text{ increases.} \quad K_c = \frac{[\text{products}]}{[\text{reactants}]}$$

10.6. The equilibrium constant for the reaction of H_2O and CO_2 to form H_2CO_3 is 1.5×10^{-3} . What is the concentration of H_2CO_3 if the concentration of H_2O is 55.5 M and the concentration of CO_2 is 0.033 M ?

- 1.5×10^{-3}
- 55.5 M
- 0.033 M
- $1.5 \times 10^{-3} \text{ M}$

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Wolfgang Guggemos

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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