

NAME: \_\_\_\_\_

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1. What amount of oxygen,  $O_2$ , (in moles) contains  $1.8 \times 10^{22}$  molecules?

☐ A 0.030

☐ B 3.0

☐ C 0.30

☐ D 0.0030

2. Which compound has the empirical formula with the greatest mass?

☐ A  $C_6H_6$

☐ B  $C_5H_{10}$

☐ C  $C_4H_{10}$

☐ D  $C_2H_6$

3.  $\_\_C2H2(g) + \_\_O2(g) \rightarrow \_\_CO2(g) + \_\_H2O(g)$

When the equation above is balanced, what is the coefficient for oxygen?

☐ A 2

☐ B 3

☐ C 5

☐ D 4

4. 3.0 dm<sup>3</sup> of sulfur dioxide is reacted with 2.0 dm<sup>3</sup> of oxygen according to the equation below.



What volume of sulfur trioxide (in dm<sup>3</sup>) is formed? (Assume the reaction goes to completion and all gases are measured at the same temperature and pressure.)

☐ A 5.0

☐ B 3.0

☐ C 2.0

☐ D 4.0

5. What will happen to the volume of a fixed mass of gas when its pressure and temperature (in Kelvin) are both doubled?

☐ A It will not change.

☐ B The change cannot be predicted.

☐ C It will increase.

☐ D It will decrease.

6. What amount (in moles) is present in 2.0 g of sodium hydroxide, NaOH?

# Grade Chemistry Questions And Answers

**Daniela Niemeyer**



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