

# Journal Entry

Date	Accounts and Explanation	Debit	Credit
(a) March 3	Merchandise inventory	\$1,150	
	Accounts payable - GW		\$1,150
	To record the purchase of merchandise on account		
(b) March 4	Merchandise inventory	75	
	Cash		75
	To record the payment of shipping charges.		
(c) March 5	Accounts payable - GW	150	
	Merchandise inventory		150
	To record the return of merchandise on account for purchases on March 3.		
(d) March 18	Accounts payable - GW	1,000	
	Merchandise inventory		20
	Cash		980
	To record the payment of merchandise purchased on account minus discount.		
(e) March 19	Merchandise inventory	425	
	Accounts payable - PF		425
	To record the payment of merchandise purchased on account minus discount.		
(f) March 21	Accounts payable - PF	25	
	Merchandise inventory		25
	To record the receipt of allowance from the vendor from the purchase on March 19.		
(g) March 29	Accounts payable - PF	\$400	
	Merchandise inventory		\$4
	Cash		\$396
	To record the payment of merchandise purchased on account minus discount.		

# Fundamental Accounting Principles Serial Problem 5

## Answers

**JE Gale**



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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is  $2\text{Na}^+\text{aq} + 2\text{Cl}^-\text{aq} + \text{Cu}^{2+}\text{aq} + \text{SO}_4^{2-}\text{aq} \rightarrow 2\text{Na}^+\text{aq} + \text{SO}_4^{2-}\text{aq} + \text{CuCl}_2\text{s}$  the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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net ionic equation  $\text{H}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{H}_2\text{O}\text{l} + \text{CO}_2\text{g}$  here is a video on spectator ions  
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web the six  $\text{Na}^+\text{aq}$  ions and the six  $\text{Cl}^-\text{aq}$  ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $3\text{Ba}^{2+}\text{aq} + 2\text{PO}_4^{3-}\text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2\text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3\text{aq} + \text{KCl}\text{aq} \rightarrow \text{AgCl}\text{s} + \text{KNO}_3\text{aq}$  total ionic  $\text{Ag}^+\text{aq} + \text{NO}_3^-\text{aq} + \text{K}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl}\text{s} + \text{K}^+\text{aq} + \text{NO}_3^-\text{aq}$  net ionic  $\text{Ag}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl}\text{s}$  2

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^+\text{aq} + \text{OH}^-\text{aq} \rightarrow \text{H}_2\text{O}\text{l}$  the spectator ions in  $\text{KOH}\text{aq}$  and  $\text{HNO}_3\text{aq}$  are  $\text{K}^+\text{aq}$  and  $\text{NO}_3^-\text{aq}$  how to identify and list the

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anions an ion with a net charge of zero is electrically neutral

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web the ag aq and cl aq ions become agcl s but the na aq ions and the no 3 aq ions stay as na aq ions and no 3 aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be 2cl aq cu 2 aq cucl 2 s 2cl aq cu 2 aq cucl 2 s these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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