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# Fema Ics 80study Guide

**L Cohen**

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$$\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$$

$$3\text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 6\text{NaBr} \rightarrow 3\text{SnO}_2 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$$

using the following equation

$$2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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$$\text{br} + \text{f}_2 \rightarrow \text{br}_2$$

$$\text{c}_6\text{h}_{10} + \text{o}_2 \rightarrow \text{co}_2 + \text{h}_2\text{o}$$

$$\text{br} + \text{hbr} \rightarrow \text{khco}_3 + \text{h}_2\text{o} + \text{kbr} + \text{co}_2$$

$$\text{ga}_2\text{so}_3 + \text{nabr} \rightarrow \text{sno}_2 + \text{nf}_3 + \text{snf}_2 + \text{n}_2\text{o}_3$$

$$\text{naoh} + \text{h}_2\text{so}_4 \rightarrow \text{h}_2\text{o} + \text{na}_2\text{so}_4$$

the stoichiometry of a balanced chemical equation identifies the maximum amount of product that can be obtained the stoichiometry of a reaction describes the relative amounts of reactants and products in a balanced chemical equation

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equations 1 2 3 4 5 i n2 3 f2 2 nf3 2 cah 40 17 02 12 co2 10 h2o hbr khco3 h2o kbr co2 gabr na2so3 gaz so3 3 nabr sno nf3 snf2

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