

Chemistry Revision

Moles, Empirical formula, Ar and Mr

A_r & M_r

- Relative atomic mass:
 - the weighted average mass of the naturally occurring isotopes of an element as compared to $1/12$ of the mass of carbon-12
 - weighted average means it factors in the abundance of each isotope
 - $(\% \times A_r) + (\% \times A_r) / 100$
- Relative formula mass:
 - the sum of all the atomic masses
 - use relative formula mass to work out percentage composition

Question: find percentage by mass of copper in copper(II) oxide, CuO

A_r : O = 16; Cu = 64

$$\begin{aligned}M_r &= 64 + 16 \\&= 80\end{aligned}$$

$$\begin{aligned}\text{Percentage of copper} &= 64/80 \times 100 \\&= 80\%\end{aligned}$$

Moles

- The number of carbon atoms in exactly 12g of carbon-12
- Formula:
 - number of moles = mass (g) / M_r
 - number of moles = number of particles / 6.02×10^{23}
- Reacting mass calculations:
 - mole (convert mass into moles)
 - mole ratio (from balanced equation)
 - un-mole (convert back to grams)
 - when un-moling in gas calculations, always multiply by 24 instead of M_r

Question: find the mass aluminium oxide when 135g of aluminium is burned with oxygen



- mole = mass/ M_r = $135/27$ = 5 mol Al
- mole ratio = $4:2$ = $2:1$ = 2.5 mol Al_2O_3
- un-mole = 2.5×102 = 255g Al_2O_3

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Wenbin Ji



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Table of Contents Igcse Edexcel Chemistry May 2013

1. Understanding the eBook Igcse Edexcel Chemistry May 2013
 - The Rise of Digital Reading Igcse Edexcel Chemistry May 2013
 - Advantages of eBooks Over Traditional Books
2. Identifying Igcse Edexcel Chemistry May 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Igcse Edexcel Chemistry May 2013
 - User-Friendly Interface
4. Exploring eBook Recommendations from Igcse Edexcel Chemistry May 2013
 - Personalized Recommendations
 - Igcse Edexcel Chemistry May 2013 User Reviews and Ratings
 - Igcse Edexcel Chemistry May 2013 and Bestseller Lists
5. Accessing Igcse Edexcel Chemistry May 2013 Free and Paid eBooks
 - Igcse Edexcel Chemistry May 2013 Public Domain eBooks
 - Igcse Edexcel Chemistry May 2013 eBook Subscription Services

- Igcse Edexcel Chemistry May 2013 Budget-Friendly Options
- 6. Navigating Igcse Edexcel Chemistry May 2013 eBook Formats
 - ePub, PDF, MOBI, and More
 - Igcse Edexcel Chemistry May 2013 Compatibility with Devices
 - Igcse Edexcel Chemistry May 2013 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Igcse Edexcel Chemistry May 2013
 - Highlighting and Note-Taking Igcse Edexcel Chemistry May 2013
 - Interactive Elements Igcse Edexcel Chemistry May 2013
- 8. Staying Engaged with Igcse Edexcel Chemistry May 2013
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Igcse Edexcel Chemistry May 2013
- 9. Balancing eBooks and Physical Books Igcse Edexcel Chemistry May 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Igcse Edexcel Chemistry May 2013
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Igcse Edexcel Chemistry May 2013
 - Setting Reading Goals Igcse Edexcel Chemistry May 2013
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Igcse Edexcel Chemistry May 2013
 - Fact-Checking eBook Content of Igcse Edexcel Chemistry May 2013
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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