

Ideal Gas Law

Chapter 9

1. A gas at a pressure of 1 atm and a volume of 10 liters has a temperature of 25°C. What is the number of moles of the gas present?
2. A sample of gas has a volume of 5 liters, a pressure of 2 atm, and a temperature of 300 K. What is the number of moles of the gas present?
3. A gas has a pressure of 1.5 atm, a volume of 20 liters, and a temperature of 25°C. What is the number of moles of the gas present?
4. A gas at a volume of 30 liters and a temperature of 27°C has a pressure of 2 atm. What is the number of moles of the gas present?
5. A gas at a pressure of 1 atm, a volume of 25 liters, and a temperature of 0°C has a number of moles equal to 2. What is the pressure of the gas at a volume of 15 liters and a temperature of 27°C?
6. A gas at a pressure of 2 atm, a volume of 10 liters, and a temperature of 25°C has a number of moles equal to 4. What is the pressure of the gas at a volume of 20 liters and a temperature of 27°C?
7. A gas at a volume of 5 liters and a temperature of 27°C has a pressure of 1 atm. What is the number of moles of the gas present?
8. A gas at a pressure of 3 atm, a volume of 20 liters, and a temperature of 25°C has a number of moles equal to 5. What is the pressure of the gas at a volume of 10 liters and a temperature of 27°C?
9. A sample of gas has a volume of 15 liters, a pressure of 1.5 atm, and a temperature of 300 K. What is the number of moles of the gas present?
10. A gas at a volume of 10 liters and a temperature of 27°C has a pressure of 2 atm. What is the number of moles of the gas present?
11. Note: These questions are meant to be examples to guide you in creating your own worksheet. The specific answers to these questions may vary based on your textbook or course material.

Ideal Gas Law Questions And Answers

Hussin A.Rothana



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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the

copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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