

## DATA SHEET

Avogadro constant, $N_A$ .....	$6.022 \times 10^{23} \text{ mol}^{-1}$
Volume of 1 mole ideal gas: at 100 kPa and	
at 0°C (273.15 K) .....	22.71 L
at 25°C (298.15 K) .....	24.79 L
Ionisation constant for water at 25°C (298.15 K), $K_w$ .....	$1.0 \times 10^{-14}$
Specific heat capacity of water .....	$4.18 \times 10^3 \text{ J kg}^{-1} \text{ K}^{-1}$

## Some useful formulae

$$\text{pH} = -\log_{10}[\text{H}^+]$$

$$\Delta H = -m C \Delta T$$

## Some standard potentials

$\text{K}^+ + \text{e}^-$	$\text{vs}$	$\text{K(s)}$	-2.94 V
$\text{Ba}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Ba(s)}$	-2.91 V
$\text{Ca}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Ca(s)}$	-2.87 V
$\text{Na}^+ + \text{e}^-$	$\text{vs}$	$\text{Na(s)}$	-2.71 V
$\text{Mg}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Mg(s)}$	-2.36 V
$\text{Al}^{3+} + 3\text{e}^-$	$\text{vs}$	$\text{Al(s)}$	-1.68 V
$\text{Mn}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Mn(s)}$	-1.18 V
$\text{H}_2\text{O} + \text{e}^-$	$\text{vs}$	$\frac{1}{2}\text{H}_2(\text{g}) + \text{OH}^-$	-0.83 V
$\text{Zn}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Zn(s)}$	-0.76 V
$\text{Fe}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Fe(s)}$	-0.44 V
$\text{Ni}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Ni(s)}$	-0.24 V
$\text{Sn}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Sn(s)}$	-0.14 V
$\text{Pb}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Pb(s)}$	-0.13 V
$\text{H}^+ + \text{e}^-$	$\text{vs}$	$\frac{1}{2}\text{H}_2(\text{g})$	0.00 V
$\text{SO}_4^{2-} + 4\text{H}^+ + 2\text{e}^-$	$\text{vs}$	$\text{SO}_2(\text{aq}) + 2\text{H}_2\text{O}$	0.16 V
$\text{Cu}^{2+} + 2\text{e}^-$	$\text{vs}$	$\text{Cu(s)}$	0.34 V
$\frac{1}{2}\text{O}_2(\text{g}) + \text{H}_2\text{O} + 2\text{e}^-$	$\text{vs}$	$2\text{OH}^-$	0.40 V
$\text{Cu}^+ + \text{e}^-$	$\text{vs}$	$\text{Cu(s)}$	0.52 V
$\frac{1}{2}\text{I}_2(\text{s}) + \text{e}^-$	$\text{vs}$	$\text{I}^-$	0.54 V
$\frac{1}{2}\text{I}_2(\text{aq}) + \text{e}^-$	$\text{vs}$	$\text{I}^-$	0.62 V
$\text{Fe}^{3+} + \text{e}^-$	$\text{vs}$	$\text{Fe}^{2+}$	0.77 V
$\text{Ag}^+ + \text{e}^-$	$\text{vs}$	$\text{Ag(s)}$	0.80 V
$\frac{1}{2}\text{Br}_2(\text{l}) + \text{e}^-$	$\text{vs}$	$\text{Br}^-$	1.08 V
$\frac{1}{2}\text{Br}_2(\text{aq}) + \text{e}^-$	$\text{vs}$	$\text{Br}^-$	1.10 V
$\frac{1}{2}\text{O}_2(\text{g}) + 2\text{H}^+ + 2\text{e}^-$	$\text{vs}$	$\text{H}_2\text{O}$	1.23 V
$\frac{1}{2}\text{Cl}_2(\text{g}) + \text{e}^-$	$\text{vs}$	$\text{Cl}^-$	1.36 V
$\frac{1}{2}\text{Cr}_2\text{O}_7^{2-} + 7\text{H}^+ + 3\text{e}^-$	$\text{vs}$	$\text{Cr}^{3+} + \frac{7}{2}\text{H}_2\text{O}$	1.36 V
$\frac{1}{2}\text{Cl}_2(\text{aq}) + \text{e}^-$	$\text{vs}$	$\text{Cl}^-$	1.40 V
$\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^-$	$\text{vs}$	$\text{Mn}^{2+} + 4\text{H}_2\text{O}$	1.51 V
$\frac{1}{2}\text{F}_2(\text{g}) + \text{e}^-$	$\text{vs}$	$\text{F}^-$	2.89 V

# Hsc Chemistry Data Sheet

**Hisao Ohtake, Satoshi Tsuneda**



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## **Table of Contents Hsc Chemistry Data Sheet**

1. Understanding the eBook Hsc Chemistry Data Sheet
  - The Rise of Digital Reading Hsc Chemistry Data Sheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry Data Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Chemistry Data Sheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry Data Sheet
  - Personalized Recommendations
  - Hsc Chemistry Data Sheet User Reviews and Ratings

- Hsc Chemistry Data Sheet and Bestseller Lists
- 5. Accessing Hsc Chemistry Data Sheet Free and Paid eBooks
  - Hsc Chemistry Data Sheet Public Domain eBooks
  - Hsc Chemistry Data Sheet eBook Subscription Services
  - Hsc Chemistry Data Sheet Budget-Friendly Options
- 6. Navigating Hsc Chemistry Data Sheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Hsc Chemistry Data Sheet Compatibility with Devices
  - Hsc Chemistry Data Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc Chemistry Data Sheet
  - Highlighting and Note-Taking Hsc Chemistry Data Sheet
  - Interactive Elements Hsc Chemistry Data Sheet
- 8. Staying Engaged with Hsc Chemistry Data Sheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc Chemistry Data Sheet
- 9. Balancing eBooks and Physical Books Hsc Chemistry Data Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc Chemistry Data Sheet
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Chemistry Data Sheet
  - Setting Reading Goals Hsc Chemistry Data Sheet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Chemistry Data Sheet
  - Fact-Checking eBook Content of Hsc Chemistry Data Sheet
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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