

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?



$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mole Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$= \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mole Cu}}{2 \text{ mol AgNO}_3} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} =$$

$$\underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol NaCl}} \times \frac{71 \text{ g Cl}_2}{1 \text{ mol Cl}_2}$$

$$= \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mol Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{0.21 \text{ moles H}_2}$$

Holt Chemistry Problem Solving Work Answers

Stoichiometry

Tao Wei



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Table of Contents **Holt Chemistry Problem Solving Work Answers Stoichiometry**

1. Understanding the eBook Holt Chemistry Problem Solving Work Answers Stoichiometry
 - The Rise of Digital Reading Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Problem Solving Work Answers Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Personalized Recommendations
 - Holt Chemistry Problem Solving Work Answers Stoichiometry User Reviews and Ratings

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- 5. Accessing Holt Chemistry Problem Solving Work Answers Stoichiometry Free and Paid eBooks
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 - Holt Chemistry Problem Solving Work Answers Stoichiometry Compatibility with Devices
 - Holt Chemistry Problem Solving Work Answers Stoichiometry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Highlighting and Note-Taking Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Interactive Elements Holt Chemistry Problem Solving Work Answers Stoichiometry
- 8. Staying Engaged with Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Problem Solving Work Answers Stoichiometry
- 9. Balancing eBooks and Physical Books Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Problem Solving Work Answers Stoichiometry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Setting Reading Goals Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Fact-Checking eBook Content of Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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