

1. Consider the following reaction:  $N_2(g) + 3H_2(g) \longrightarrow 2NH_3(g)$ , if the rate of the loss of hydrogen gas is  $0.09 \text{ mol/l}\cdot\text{s}$ , what is the rate of production of ammonia?

2. The rate law of a given reaction is  $r = kC^2$ , with  $k = 2.64 \times 10^{-3} \text{ M}^{-1} \text{ min}^{-1}$ , If the initial concentration is  $0.025 \text{ M}$ , what is the initial rate?

3. the reaction  $2NO(g) + Cl_2(g) \longrightarrow 2NOCl(g)$ , with rate law,  $r = kC_{NO}^2C_{Cl_2}$  is \_\_\_\_\_ order with respect to NO, \_\_\_\_\_ order with respect to  $Cl_2$  and \_\_\_\_\_ order overall.

4. The rate of a particular reaction is quadrupled when the temperature was increased from  $55^\circ\text{C}$  to  $60^\circ\text{C}$ . What is activation energy (in J/mol) for this process.

5. A common rule of thumb is that near room temperature, the rate of a reaction will double with each  $10^\circ\text{C}$  rise in temperature. Estimate the activation energy (in J/mol) needed for this rule to hold true.

6. The mechanism of a reaction is shown below:



- (A) What is the overall reaction?  
(B) Which compounds are intermediates?  
(C) Predict the rate law based on this mechanism.  
(D) What is the overall order of the reaction?

7. consider the following reaction,  $N_2H_4 + 2H_2O_2 \longrightarrow N_2 + 4H_2O$ . In 55 seconds,  $0.165 \text{ mol}$  of  $H_2O_2$  is consumed. The rate of production of  $N_2$  is

- (A)  $1.5 \times 10^{-3} \text{ mol}$  in 55 seconds  
(B)  $0.33 \text{ mol/s}$   
(C)  $1.5 \times 10^{-3} \text{ mol/s}$   
(D)  $0.006 \text{ mol/s}$

8. The dimensions of rate constant for reaction  $3A \longrightarrow B$  are  $(\text{l/gm mole})/\text{min}$ . Therefore the reaction order is

- (A) 0  
(B) 1  
(C) 2  
(D) 3

# Holt Chemistry Reaction Rates Answer Key

**RJ Alexander**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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