Section 7.2: Mass Relationships in Chemical Reactions

Tutorial 1 Practice, page 323-324

1. (a) Given: $m_{CoCO_1} = 0.50 \text{ g}$

Required: mass of hydrochloric acid, m_{BCI}

Solution:

Step 1. List the given value, the required value, and the corresponding molar masses.

 $CaCO_3(s) + 2HCl(aq) \rightarrow CO_2(g) + H_2O(l) + CaCl_2(aq)$

0.50 g M_{HCI}

100.09 g/mol 36.46 g/mol

Step 2. Convert mass of calcium carbonate to amount of calcium carbonate.

$$n_{\text{CaCO}_3} = 0.50 \, \text{g} \times \frac{1 \, \text{mol}_{\text{CaCO}_3}}{100.09 \, \text{g}}$$

 $n_{\text{CaCO}_2} = 0.004 \text{ 996 mol [two extra digits carried]}$

Step 3. Convert amount of calcium carbonate to amount of hydrochloric acid.

$$n_{\text{HCI}} = 0.004 996 \text{ mol}_{\text{CsCO}_{1}} \times \frac{2 \text{ mol}_{\text{HCI}}}{1 \text{ mol}_{\text{CsCO}_{1}}}$$

 $n_{\text{HC}} = 0.009 992 \text{ mol [two extra digits carried]}$

Step 4. Convert amount of hydrochloric acid to mass of hydrochloric acid.

$$m_{\text{HCI}} = (0.009 \ 992 \ \text{pert}) \left(\frac{36.46 \ \text{g}}{1 \ \text{John}} \right)$$

$$m_{\text{HCI}} = 0.36 \text{ g}$$

Statement: 0.50 g of calcium carbonate will neutralize 0.36 g of hydrochloric acid.

(b) Given: $m_{CaCO_3} = 0.50 \text{ g}$; $n_{CaCO_4} = 0.004 996 \text{ mol}$

Required: mass of calcium chloride, mcaci-

Solution:

Step 1. List the given values, the required values, and the corresponding molar masses.

 $CaCO_3(s) + 2 HCl(aq) \rightarrow CO_2(g) + H_2O(l) + CaCl_2(aq)$ 0.004 996 mol n_{CaCl_3}

0.004 996 mot 0.50 g

Michella

100.09 g/mo

110.98 g/mol

Step 2. Convert amount of calcium carbonate to amount of calcium chloride.

$$n_{\text{CoO}_2} = n_{\text{CoO}_2}$$

$$n_{\text{CaCI}_c} = 0.004 996 \text{ mol}$$

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Caren Van Slyke

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