2

SECTION A

Answer ALL questions in this Section. Write your answers in the spaces provided.

- (a) The pH of human blood is maintained within a narrow range from 7.35 to 7.45 by a natural buffer system consisting of carbonic acid, H₂CO₂(aq), and hydrogenearbonate ions, HCO₃⁻ (aq).
 - A buffer solution containing H₂CO₂(aq) and HCO₃⁻ (aq) in equal concentrations has a pH of 6.10. Calculate the dissociation constant, K₈, for H₂CO₂(aq).

$$pH = pK_a + log \frac{[HCO_3^-]}{[H_1CO_1]}$$

In the solution, $[HCO_3] = [H_2CO_3]$

$$K_a = 7.94 \times 10^{-7} \text{ mol dm}^{-3}$$

(ii) Calculate the ratio of concentrations of HCO₃" (aq) and H₂CO₃(aq) in blood at pH 7.40.

$$7.4 = -\log(7.94 \times 10^{-7}) + \log \frac{[HCO_3^{-1}]}{[H_2CO_3]}$$

$$log \frac{[HCO_3^-]}{[H_2CO_3]} = 1.3$$

 $\therefore \frac{[HCO_3^-]}{[H_2CO_3]} = 20$

(Accept answers from 19.8 to 20.0) (1 mark for method; 1 mark for answer)

(iii) (I) Would the blood pH of a person increase or decrease in the course of physical exertion? Explain your answer.

> During physical exertion, the concentration of CO₃ in blood increases. 16 The equilibrium

$$H_2CO_3 = H_2O \Rightarrow HCO_3^- + H_2O^+$$
 % shifts to the right. ... pH of blood will drop. %, %

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The State of Education Social Policy and Development Centre, 2003

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