#### 15 Hydrocarbons

#### Alkanes

#### Chemical Reactivity

Alkanes do not react with many reagents.

This is because the C-C bond and the C-H bond are relatively strong and so cannot form stronger bonds with other elements that easily and there are no polar side groups .

There main reactions are combustion reactions in air although they can react with halogens in the presence of UV light

#### Combustion

#### Complete Combustion

In excess oxygen alkanes will burn with complete combustion

The products of complete combustion are CO, and H,O.

 $C_2H_2(g) + 3.5 O_3(g) \rightarrow 2CO_3(g) + 3 H_3O(l)$ 

Alkanes readily burn in the presence of oxygen. This combustion of alkanes is highly exothermic, explaining their use as fuels.

#### Reaction of alkanes with bromine / chlorine in UV light

#### Synthesis of chloroalkanes

In the presence of UV light alkanes react with chlorine to form a mixture of products with the halogens substituting hydrogen atoms.



This is the overall reaction, but a more complex mixture of products is actually formed

To understand this reaction fully we must look in detail at how it proceeds step by step. This is called its mechanism

The MECHANISM for this reaction is called a FREE RADICAL SUBSTITUTION

It proceeds via a series of steps: STEP ONE: Initiation STEP TWO: Propagation STEP THREE: Termination

#### STEP ONE Initiation

Essential condition: UV light

Cl<sub>2</sub> → 2Cl<sup>-</sup>

: CI: CI: + × CI:

The UV light supplies the energy to break the CI-CI bond. It is broken in preference to the others as it is the weakest.

The bond has broken in a process called homolytic fission.

When a bond breaks by homolytic fission it forms Free Radicals.

Free Radicals do not have a charge and are represented by a

## each atom gets one electron from the covalent bond

A Free Radical is a reactive species which possess an unpaired electron

#### **STEP TWO Propagation**

-CH<sup>2</sup> + Cl<sup>5</sup> → CH<sup>2</sup>Cl + Cl.

The chlorine free radicals are very reactive and remove an H from the methane leaving a methyl free radical

The methyl free radical reacts with a Cl<sub>2</sub> molecule to produce the main product and another Cl free radical

All propagation steps have a free radical in the reactants and in the products.

As the CI free radical is regenerated, it can react with several more alkane molecules in a CHAIN REACTION

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