

HOLT MODERN CHEMISTRY -- CHAPTER 3 REVIEW

*The following pages contain the bulk (but not all) of the information for the chapter 3 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test. ***** Test items will be recall, examples, and/or application of this content. ******

OUTCOMES

- 3.1: Explain law of conservation of mass (*T & R*)
- 3.1-3.2: Explain how the atom changed over time (Rutherford - Bohr models) (*F & PK*)
- 3.3: Contrast the nature and properties of the three subatomic particles: proton, neutron, and electron (*T & R*)
- 3.3: Use atomic number and mass number to find the number of subatomic particles of any given atom (*F & PK*)
- 3.3: Draw contrasting Bohr models of isotopes (*F & PK*)
- 3.3: Define mole and Avogadro's number (*T & R*)

3.1 THE ATOM: FROM PHILOSOPHICAL IDEA TO SCIENTIFIC THEORY

- **Vocabulary**
 - **law of conservation of mass** -- the law that states that mass cannot be created or destroyed in ordinary chemical and physical changes
 - **law of definite proportions** -- the law that states that a chemical compound always contains the same elements in exactly the same proportions by weight or mass
 - **law of multiple proportions** -- the law that states that when two elements combine to form two or more compounds, the mass of one element that combines with a given mass of the other is in the ratio of small whole numbers
- **Chapter Highlights**
 - The idea of atoms has been around since the time of the ancient Greeks. In the nineteenth century, John Dalton proposed a scientific theory of atoms that can still be used to explain properties of most chemicals today.
 - Matter and its mass cannot be created or destroyed in chemical reactions.
 - The mass ratios of the elements that make up a given compound are always the same, regardless of how much of the compound there is or how it was formed.
 - If two or more different compounds are composed of the same two elements, then the ratio of the masses of the second element combined with a certain mass of the first element can be expressed as a ratio of small whole numbers.

3.2 THE STRUCTURE OF THE ATOM

- **Vocabulary**
 - **atom** -- the smallest unit of an element that maintains the chemical properties of that element
 - **nuclear forces** -- the interaction that binds protons and neutrons, protons and protons, and neutrons and neutrons together in a nucleus
- **Chapter Highlights**
 - Cathode-ray tubes supplied evidence of the existence of electrons, which are negatively charged subatomic particles that have relatively little mass.
 - Rutherford found evidence for the existence of the atomic nucleus by bombarding gold foil with a beam of positively charged particles.
 - Atomic nuclei are composed of protons, which have an electric charge of +1, and (in all but one case) neutrons, which have no electric charge.
 - Atomic nuclei have radii of about 0.001 pm (pm = picometers; $1 \text{ pm} \times 10^{-12} \text{ m}$), and atoms have radii of about $40\text{--}270 \text{ pm}$.

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W/Corr Sci 2001 Holt Rinehart & Winston, 2002-03 *Biology*, 1998 **Basic Chemistry for Water and Wastewater Operators** Darshan Singh Sarai, 2005-06 **Modern Physical Chemistry** G.H. Duffey, 2013-11-11 In this new textbook on physical chemistry fundamentals are introduced simply yet in more depth than is common Topics are arranged in a progressive pattern with simpler theory early and more complicated theory later General principles are induced from key experimental results Some mathematical background is supplied where it would be helpful Each chapter includes worked out examples and numerous references Extensive problems review and discussion questions are included for each chapter More detail than is common is devoted to the nature of work and heat and how they differ Introductory Caratheodory theory and the standard integrating factor for dGrev are carefully developed The fundamental role played by uncertainty and symmetry in quantum mechanics is emphasized In chemical kinetics various methods for determined rate laws are presented The key mechanisms are detailed Considerable statistical mechanics and reaction rate theory are then surveyed Professor Duffey has given us a most readable easily followed text in physical chemistry **Aircraft Performance Engineering for Reciprocating Engines** United States. Department of the Air Force, 1954 **The Scientific Imagination** Gerald James Holton, 1998 Using firsthand accounts gleaned from notebooks interviews and correspondence of such twentieth century scientists as Einstein Fermi and Millikan Holton shows how the idea of the scientific imagination has practical implications for the history and philosophy of science and the larger understanding of the place of science in our culture **Molecular Physics and Elements of Quantum Chemistry** Hermann Haken, Hans Christoph Wolf, 2013-04-18 Since the publication of the first edition of this book there have been many important new developments in the field of molecular physics The new methods and results which are most significant for students are treated extensively in this second edition Among these are in particular single molecule spectroscopy and the field of molecular electronics which is in a stage of rapid development including the areas of electroluminescence and organic light emitting diodes In addition we have extended and corrected the earlier material in a number of places We have also included exercises in this new edition they will allow students to deepen their understanding and offer a basis for further individual study The complete solutions to the exercises can be found on the Internet under www.springeronline.com 3 540 40792 S We are grateful to Mr C D Bachem and Dr Th Schneider of the Springer Verlag for their continuous and very agreeable cooperation during the preparation of the book We thank our colleague Prof W D Brewer for his competent translation Stuttgart February 2004 H Haken H C Wolf Preface to the First Edition This textbook is intended for use by students of physics physical chemistry and theoretical chemistry The reader is presumed to have a basic knowledge of atomic and quantum physics at the level provided for example by the first few chapters in our book *The Physics of Atoms and Quanta* *An Introduction to Atmospheric Physics* Robert G. Fleagle, Joost A.

Businger,1981-01-09 This book is addressed to those who wish to understand the relationship between atmospheric phenomena and the nature of matter as expressed in the principles of physics The interesting atmospheric phenomena are more than applications of gravitation of thermodynamics of hydrodynamics or of electrodynamics and mastery of the results of controlled experiment and of the related theory alone does not imply an understanding of atmospheric phenomena This distinction arises because the extent and the complexity of the atmosphere permit effects and interactions that are entirely negligible in the laboratory or are deliberately excluded from it the objective of laboratory physics is by isolating the relevant variables to reveal the fundamental properties of matter whereas the objective of atmospheric physics or of any observational science is to understand those phenomena that are characteristic of the whole system For these reasons the exposition of atmospheric physics requires substantial extensions of classical physics It also requires that understanding be based on a coherent way of seeing the ensemble of atmospheric phenomena Only then is understanding likely to stimulate still more general insights

The Investigation of Organic Reactions and Their Mechanisms Howard Maskill,2008-04-15 A range of alternative mechanisms can usually be postulated for most organic chemical reactions and identification of the most likely requires detailed investigation Investigation of Organic Reactions and their Mechanisms will serve as a guide for the trained chemist who needs to characterise an organic chemical reaction and investigate its mechanism but who is not an expert in physical organic chemistry Such an investigation will lead to an understanding of which bonds are broken which are made and the order in which these processes happen This information and knowledge of the associated kinetic and thermodynamic parameters are central to the development of safe efficient and profitable industrial chemical processes and to extending the synthetic utility of new chemical reactions in chemical and pharmaceutical manufacturing and academic environments Written as a coherent account of the principal methods currently used in mechanistic investigations at a level accessible to academic researchers and graduate chemists in industry the book is highly practical in approach The contributing authors an international group of expert practitioners of the techniques covered illustrate their contributions by examples from their own research and from the relevant wider chemical literature The book covers basic aspects such as product analysis kinetics catalysis and investigation of reactive intermediates It also includes material on significant recent developments e g computational chemistry calorimetry and electrochemistry in addition to topics of high current industrial relevance e g reactions in multiphase systems and synthetically useful reactions involving free radicals and catalysis by organometallic compounds

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