Equilibrium: Assessment Review

Management

Skill - The student will be able to draw and explain how a catalyst changes a reaction rate graph and discuss other changes that disrupt an equilibrium system.

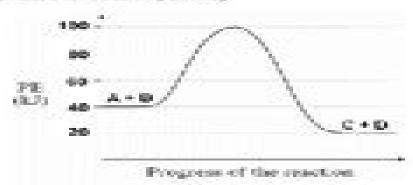
- Using the reaction diagram to the right, representing the reaction: A + B → C + D dram a new curve that shows how a catalyst would affect the reaction pathway.
- Place a checkmark next to the actions that would most likely <u>secrepse</u> the reaction rate:

Lowering the temp of the system

Increasing the concentration of substances present.
Gending a solid in to fine particles.

Removing a enzyme catalyst.

 Explain why all reactions have activation energy, using your knowledge of collision model.



- Evglain why different reactions have different activation energies.
- Describe how the activation energy of a reaction affects the overall rate of the chemical reaction differently than a change in concentration, pressure, or temp.
- A rule of thumb used by most chemistic is that the rate of a chemical reaction can be doubled by increasing the reaction temperature by ten degrees Celsius. Explain this drastic increase in reaction rate using your knowledge of the collegion model.

Skill - The student will be able to write the equilibrium expression, given a reaction. When a reaction from reaction expression reaction from the foreign reaction and the same rate that the foreign reaction products and reaction from the same had account of products and reaction for not thank. This product is called dynamic equilibrium because the account of products and reaction for not thanks. This product is called dynamics equilibrium.

A reaction that has reached equilibrium can be described with a ratio known as the equilibrium expression. This expression relates the concentration of products divided by the concentration of reactions, solving for the equilibrium constitutes constituted the product of products and partners are included.

Balance the following equations & write the equilibrium expressions needed to solve for E:

- T. Outigh # Outight
- 8. HCl (g) + O₂ (g) 0 H₂O (l) + O₃ (g);
- 9. 55 (sil + H,51g) \$ 55,5,160 + H,1g)

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