

# Equilibrium

At what point is a reversible reaction "completed"?

## Why?

Most of the reactions that we have studied this year have been "forward" reactions—once the reactant has changed into the product it stays that way. We can assume that eventually the reaction will "finish" when the limiting reactant runs out. However, in a reversible reaction (one that can take place in both directions), once the product is formed it can turn back into the reactant and a continuous cycle occurs. When do reactions such as these "finish"? Is there ever an end, and can we predict how much reactant or product will be present when the reaction is "done?"

## Model 1 – A Reversible Reaction



1. What is the reactant in the reaction in Model 1?
2. What is the product in the reaction in Model 1?
3. What is the significance of the double arrow in the equation in Model 1?



4. Imagine that the reaction in Model 1 starts with 100 molecules of A and zero molecules of B. Explain why the concentration of substance A will never reach zero.
5. Imagine the reaction in Model 1 starts with zero molecules of A and 100 molecules of B. Will the concentration of substance B increase or decrease as the reaction proceeds? Explain.

# Equilibrium Pogil Answers

**Tobias Bleicker**



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