

Step I: Calculation of K_p for the reaction

Let the total mass of the gaseous mixture = 100 g

Mass of CO in the mixture = 90.55 g

Mass of CO₂ in the mixture = (100 - 90.55) = 9.45 g

$$\text{No. of moles of CO} = \frac{90.55 \text{ g}}{(28 \text{ g mol}^{-1})} = 3.234 \text{ mol}$$

$$\text{No. of moles of CO}_2 = \frac{9.45}{(44 \text{ g mol}^{-1})} = 0.215 \text{ mol}$$

$$p_{\text{CO}} \text{ in the mixture} = \frac{(3.234 \text{ mol})}{(3.234 + 0.215)} \times 1 \text{ atm} = \frac{(3.234 \text{ mol})}{(3.449 \text{ mol})} \times 1 \text{ atm} = 0.938 \text{ atm}$$

$$p_{\text{CO}_2} \text{ in the mixture} = \frac{(0.215 \text{ mol})}{(3.449 \text{ mol})} \times 1 \text{ atm} = 0.062 \text{ atm}$$



$$K_p = \frac{p^2_{\text{CO}}}{p_{\text{CO}_2}} = \frac{(0.938 \text{ atm})^2}{(0.062 \text{ atm})} = 14.19 \text{ atm}$$

Step II. Calculation of K_c for the reaction.

$$K_c = \frac{K_p}{(RT)^{\Delta n_g}}$$

$$K_p = 14.19 \text{ atm,}$$

$$R = 0.0821 \text{ L atm K}^{-1} \text{ mol}^{-1}, T = 1127 \text{ K}; \Delta n_g = 2 - 1 = 1$$

$$K_c = \frac{(14.19 \text{ atm})}{(0.0821 \text{ L atm K}^{-1} \text{ mol}^{-1}) \times (1127 \text{ K})^1} = 6.46$$

Equilibrium Chemistry Problems

James N. Jensen



Equilibrium Chemistry Problems:

Equilibrium Calculations John T. Donoghue, 1971 Freshman Chemistry Problems and how to Solve Them Michell J. Sienko, 1964 Chemistry Problem Solver A. Lamont Tyler, Editors of REA, 2012-04-27 Each Problem Solver is an insightful and essential study and solution guide chock full of clear concise problem solving gems All your questions can be found in one convenient source from one of the most trusted names in reference solution guides More useful more practical and more informative these study aids are the best review books and textbook companions available Nothing remotely as comprehensive or as helpful exists in their subject anywhere Perfect for undergraduate and graduate studies Here in this highly useful reference is the finest overview of chemistry currently available with hundreds of chemistry problems that cover everything from atomic theory and quantum chemistry to electrochemistry and nuclear chemistry Each problem is clearly solved with step by step detailed solutions DETAILS The PROBLEM SOLVERS are unique the ultimate in study guides They are ideal for helping students cope with the toughest subjects They greatly simplify study and learning tasks They enable students to come to grips with difficult problems by showing them the way step by step toward solving problems As a result they save hours of frustration and time spent on groping for answers and understanding They cover material ranging from the elementary to the advanced in each subject They work exceptionally well with any text in its field PROBLEM SOLVERS are available in 41 subjects Each PROBLEM SOLVER is prepared by supremely knowledgeable experts Most are over 1000 pages PROBLEM SOLVERS are not meant to be read cover to cover They offer whatever may be needed at a given time An excellent index helps to locate specific problems rapidly **Problems and Problem Solving in Chemistry Education** Georgios Tsaparis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner **Concepts And Problems In Physical Chemistry** P.S. Raghavan, 1997 Contents

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Aquatic Chemistry Concepts, Second Edition James F. Pankow,2019-10-31 Aquatic Chemistry Concepts Second Edition is a fully revised and updated textbook that fills the need for a comprehensive treatment of aquatic chemistry and covers the many complicated equations and principles of aquatic chemistry It presents the established science of equilibrium water chemistry using the uniquely recognizable step by step Pankow format which allows a broad and deep understanding of aquatic chemistry The text is appropriate for a wide audience including undergraduate and graduate students industry professionals consultants and regulators Every professional using water chemistry will want this text within close reach and students and professionals alike will expect to find at least one copy on their library shelves Key Features Extremely thorough one of a kind treatment of aquatic chemistry Discussions of how to carry out complex calculations regarding the chemistry of lakes rivers groundwater and seawater Numerous example problems worked in complete detail Special foreword by Jerry L Schnoor

The Gas Engineer ,1915 *The Gas Engineer's Magazine* ,1915

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