

MITSUBISHI

AC SERVO

DATE: 2015-08

MODEL **MR-J2S-40A-EG011**

POWER : 400W

INPUT : 3.2A DC283-326V

OUTPUT: 170V 0-360Hz 2.8A

SERIAL: Q58H9L057 CM

BC336U509



MITSUBISHI ELECTRIC CORPORATION

MADE IN JAPAN

PASSED

Mitsubishi Melservo Mr J2s 40a Manual

Julia Schneider



Mitsubishi Melservo Mr J2s 40a Manual:

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web 1 consider the reaction $\text{i}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$ a 80 0 grams of iodine v oxide i_2O_5 reacts with 28 0 grams of carbon monoxide co determine the mass of iodine I_2 which could be produced b if in the above situation only 0 160 moles of iodine I_2 was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 8.3 mole H₂S $\times$ 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N
 what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation C₃H₆ + 3O₂ → 3CO + 3H₂O if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
 step one figure out how many moles of the limiting reagent you have
 $21.3 \text{ grams O}_2 \div 32 \text{ grams/mole} = 0.6656 \text{ moles}$

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web percent yield is a measure of the efficiency of the experimental design
 yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 4 Al + 3 O₂ → 2 Al₂O₃
 b determine the theoretical yield of Al₂O₃
 $162.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 153.9 \text{ g Al}_2\text{O}_3$
 302.3 g Al₂O₃ / 26.982 g Al = 4 mol Al
 1 mol Al₂O₃
 c determine the percent yield
 $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

7.2 limiting reagent and reaction yields chem 1114 - Jul 02 2022

web 7 Na₂C₂O₄ is the limiting reactant percent yield = $\frac{86.6}{100} = 86.6\%$
 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless
 10 nitrogen is the limiting reagent
 11 yes methane is the limiting reagent
 12 C is the limiting reagent
 4.33 g of H₂ are left over

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web limiting reagent and percent yield practice worksheet answer key limiting reactant yield practice worksheet 1 methanol CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is reacted with 8.6 g of H_2

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