

MASS MOLE

STOICHIOMETRY WORKSHEET

1. Consider the following reaction:



- If 15 g of N_2O_4 was produced, how many moles of O_2 were required?
- If 4×10^{-2} moles of oxygen reacted, how many grams of N_2 were needed?

2. Consider the following reaction:



- What is the mass of potassium nitrate produced when 2.04 moles of potassium phosphate react?
- If 5.8 g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

3. Consider the following reaction:



- If you have 5.5 mol of CaC_2 , how much C_2H_2 do you get?
- How many moles of water are needed when 65 g of CaC_2 have reacted?

4. Hydrogen sulfide gas burns in oxygen to produce sulfur dioxide and water vapor:



What mass of oxygen gas is consumed in a reaction that produces 4.6 mol SO_2 ?

Mass Mass Stoichiometry Packet Answers

Arshad Iqbal



Mass Mass Stoichiometry Packet Answers:

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