

CHAPTER 3 TEST*Atoms: The Building Blocks of Matter*

MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice that best completes the statement or answers the question.

1. _____ The behavior of cathode rays in a glass tube containing gas at low pressure led scientists to conclude that the rays were composed of _____.
- a. energy
b. positive particles
c. negative particles
d. neutral particles
2. _____ The basic principles of atomic theory that are still recognized today were first conceived by _____.
- a. Avogadro
b. Bohr
c. Dalton
d. Rutherford
3. _____ An example of the law of multiple proportions is the existence of _____.
- a. FeCl_2 and $\text{Fe(NO}_3)_3$
b. O_2 and O_3
c. CO and CO_2
d. FeCl_2 and $\text{Fe(NO}_3)_3$
4. _____ Atoms of the same element can differ in _____.
- a. chemical properties
b. mass number
c. atomic number
d. number of protons and electrons
5. _____ Dalton's atomic theory helped to explain the law of conservation of mass because it stated that atoms _____.
- a. could not combine
b. were indivisible
c. all had the same mass
d. could not be created or destroyed
6. _____ Millikan's experiments determined _____.
- a. that the electron carried no charge
b. that the electron carried the smallest possible positive charge
c. the approximate value of the electron's mass
d. that the electron had no mass
7. _____ In Rutherford's experiment, a small percentage of the positively charged particles bombarding the metal's surface _____.
- a. were slightly deflected as they passed through the metal
b. were deflected back toward the source from the metal
c. passed straight through the metal
d. combined with the metal
8. _____ Most of the volume of an atom is made up of the _____.
- a. nucleus
b. nucleon
c. electron cloud
d. proton

Holt Mcdougal Chapter 7 Section Quiz Chemistry

Mike Jess



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