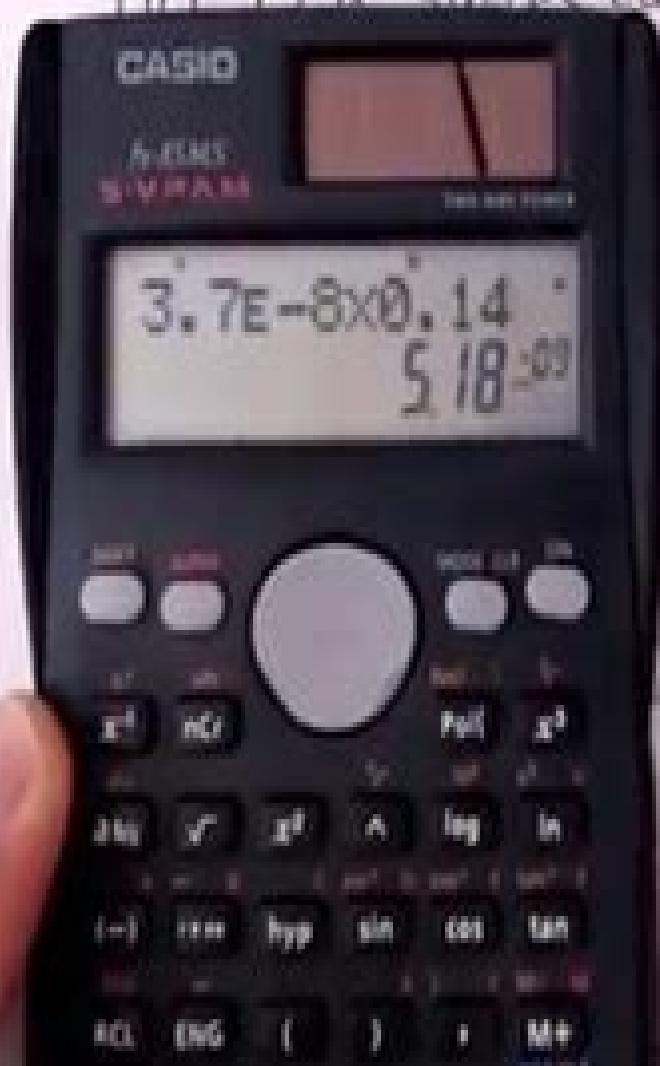


- (a) 0.14 mol dm^{-3} solutions of hydrochloric acid, HCl , and chloric(II) acid, HClO ($K_a = 1.1 \times 10^{-8}$), have different pH values.

Explain why the pH values are different and calculate the pH of 0.14 mol dm^{-3} solutions of HCl and HClO to two decimal places.

Show any working in calculations.

HCl is a strong acid and HClO is



$$\rightarrow -\log 0.14 = 0.85.$$

$$7 \times 10^{-8}$$

$$[\text{H}^+]^2 = 3.7 \times 10^{-8} \times 0.14$$

$$= 5.18 \times 10^{-9}$$

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**LeRoy Lad Panek, Mary M. Bendel-
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