# 7.05 Equilibrium Lab

Instructions: Complete all sections of this report and answer all questions in complete sentences for full credit. Use the <u>grading rubric</u> before you begin.

Click here to watch the <u>lab completion video</u>. We show the entire lab simulation (so you can record the data; don't need to be logged into the course) and explain the conclusion/post lab

Title: The Equilibrium Lab

Objective(s): How increasing or decreasing the concentration of one component of a reaction affects its equilibrium.

Hypothesis:  $\downarrow\downarrow\downarrow$ 

Complete the following statements for your hypotheses. Chaose one of the options in brackets by highlighting, bolding, or placing your answer in a different color. (Careful: the first two are about the REACTANTS changing; the last two are about the PRODUCTS changing)

- If the amount of reactant increases, the rate of the [forward or reverse] reaction will [increase] or decrease] to reach a new equilibrium.
- If the amount of reactant decreases, the rate of the [forward or reverse] reaction will [increase or decrease] to reach a new equilibrium.
- If the amount of product increases, the rate of the [forward or reverse] reaction will [increase or decrease] to reach a new equilibrium.
- If the amount of product decreases, the rate of the [forward or reverse] reaction will [increase] or decrease] to reach a new equilibrium.

## Materials:

For this lab, you are not expected to list the materials. But still notice the materials (lab equipment and chemicals) used in the lab to help you analyze the data, write the conclusion, and answer the post-lab questions.

# **Equilibrium Candy Lab Answers**

**Daniel C. Harris** 

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,1991

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