

A metallic bond is formed when 2 or more atoms exchange valence electrons to increase stability.

ionic bond is formed when valence electrons are transferred from one atom to another.

metals = atoms lose electrons to become [+] charged cations.

nonmetals gain electrons to become [-] charged anions.

In ionic compounds that ions are arranged in a crystal lattice. Strong forces hold the ions together.

Properties of ionic compounds:

- high melting and boiling points
- hard & not easily scratched
- conduct electricity when molten or dissolved because the ions are free to move.

covalent bond = electrons are shared. Forming molecules.

Covalent compounds have weak forces holding the molecules together.

Properties of covalent compounds:

- lower melting and boiling points
- Many covalent compounds are volatile liquids or gases.
- soft, easier to scratch.
- are not conductors of electricity.
- electronegativity = property that tells how strong an

Episode 501 Intro To Bonding

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