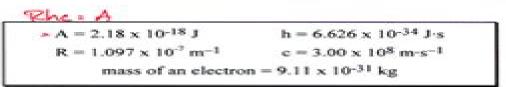
7&8 • Atomic Structure & Periodicity

PRACTICE TEST



- What wavelength corresponds to a frequency of 11. $8.22 \times 10^9 \text{ Hz?}$
 - a) 0.307 m
- d) 0.110 m
- b) 0.0365 m 500 (c) 27.4 m
- A radio station transmits at 110 MHz (110 x 106
 - Hz). What wavelength is this radio wave?
 - a) 3.65 x 10⁻⁵ m
- c) 3.81 x 10⁻⁵ m

- b) 3.30 m

- Which one of the following is NOT a proper unit for frequency?
 - a) Hz b) s⁻¹

- Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen
 - a) 0.12334 m
 - d) 4.1029 x 10⁻⁷ m
 - b) 24.373 m
- e) 36.559 m
- c) 2.7353 x 10⁻⁷ m DE = 484+1519T
- What is the relationship between the energy of a photon of light and its frequency?
 - a) E = v

e) $E = \frac{V}{h}$

- 6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level? 3 - E2
 - a) 1.52 x 10⁴ J
- d) 4.48 x 10⁻¹⁹ J
- b) 3.63 x 10⁻¹⁹ J
 - (e) 3.03 x 10⁻¹⁹ J
- c) 2.18 x 10-19 J -2.18 YOU W J
- What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.

 - a) 3.03 x 10⁻¹² m c) 3.30 x 10¹¹ m

- What resultant is expected from the interference of the two waves shown below?







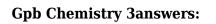
- Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?

C) ME,

d) esc.

Gpb Chemistry 3answers

Peipei Pang



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Table of Contents Gpb Chemistry 3answers

- 1. Understanding the eBook Gpb Chemistry 3answers
 - The Rise of Digital Reading Gpb Chemistry 3answers
 - o Advantages of eBooks Over Traditional Books
- 2. Identifying Gpb Chemistry 3answers
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - $\circ \ \ Determining \ Your \ Reading \ Goals$
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gpb Chemistry 3answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Gpb Chemistry 3answers
 - Personalized Recommendations
 - Gpb Chemistry 3answers User Reviews and Ratings

- Gpb Chemistry 3answers and Bestseller Lists
- 5. Accessing Gpb Chemistry 3answers Free and Paid eBooks
 - o Gpb Chemistry 3answers Public Domain eBooks
 - Gpb Chemistry 3answers eBook Subscription Services
 - Gpb Chemistry 3answers Budget-Friendly Options
- 6. Navigating Gpb Chemistry 3answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Gpb Chemistry 3answers Compatibility with Devices
 - Gpb Chemistry 3answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gpb Chemistry 3answers
 - Highlighting and Note-Taking Gpb Chemistry 3answers
 - Interactive Elements Gpb Chemistry 3answers
- 8. Staying Engaged with Gpb Chemistry 3answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gpb Chemistry 3answers
- 9. Balancing eBooks and Physical Books Gpb Chemistry 3answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gpb Chemistry 3answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gpb Chemistry 3answers
 - Setting Reading Goals Gpb Chemistry 3answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gpb Chemistry 3answers
 - Fact-Checking eBook Content of Gpb Chemistry 3answers
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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