

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, but are made of different molecules and are made up of different numbers of particles. (The particles of the gases are different sizes.)
- At standard atmospheric conditions, all the particles of the same gas have the same kinetic energy.
- At standard atmospheric conditions, the average kinetic energy of the particles of a gas is proportional to the absolute temperature.

Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three different phases: solid, liquid, and gas.



Figure 1: Representation of particles in different phases of matter

Model 1 represents the particles in different phases of matter. The particles in the solid phase are closely packed and have a regular arrangement. The particles in the liquid phase are closely packed but have a disordered arrangement. The particles in the gaseous phase are widely spaced and have a disordered arrangement.

Key Concepts

1. Model 1 shows the representation of particles in different phases of matter. The particles in the solid phase are closely packed and have a regular arrangement. The particles in the liquid phase are closely packed but have a disordered arrangement. The particles in the gaseous phase are widely spaced and have a disordered arrangement.

Table 1: Representation of particles in different phases of matter

	Solid	Liquid	Gas
How are the particles arranged?	The particles are closely packed and have a regular arrangement. The particles are held together by strong forces.	The particles are closely packed but have a disordered arrangement. The particles are held together by moderate forces.	The particles are widely spaced and have a disordered arrangement. The particles are held together by weak forces.
How do the particles move?	The particles vibrate about their fixed positions. They have a low kinetic energy.	The particles move randomly and have a moderate kinetic energy.	The particles move randomly and have a high kinetic energy.

Lw Pogil Kinetic Molecular Theory Key

William Pingry Boynton



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Applications of the Kinetic Theory to Gases, Vapors, Pure Liquids, and the Theory of Solutions William Pingry Boynton, 1904 *Molecules and the Molecular Theory of Matter* Allan Douglas Risteen, 1895 **Kinetic Theory** S. G. Brush, 2016-07-04 Kinetic Theory Volume I The Nature of Gases and of Heat deals with kinetic theory and the nature of gases and heat A comprehensive account of the life works and historical environment of a number of scientists such as Robert Boyle and Hermann von Helmholtz is presented This volume is comprised of 11 chapters and begins with an overview of the caloric theory the principle of conservation of energy the virial theorem and atomic magnitudes The discussion then turns to the qualitative atomic theory of the spring of the air proposed by Robert Boyle Isaac Newton's repulsion theory Daniel Bernoulli's theory on the properties and motions of elastic fluids especially air and George Gregory's theory on the existence of fire Subsequent chapters focus on Robert Mayer's theory on the forces of inorganic nature James Joule's theory on matter living force and heat Hermann von Helmholtz's theory on the conservation of force and Rudolf Clausius's theory on the nature of heat James Clerk Maxwell's dynamical theory of gases is also examined This book is written primarily for students and research workers in physics as well as for historians of science Irreversible Processes Stephen G. Brush, 2013-10-22 Kinetic Theory Volume 2 Irreversible Processes deals with the kinetic theory of gases and the irreversible processes they undergo It includes the two papers by James Clerk Maxwell and Ludwig Boltzmann in which the basic equations for transport processes in gases are formulated together with the first derivation of Boltzmann's H theorem and a discussion of this theorem along with the problem of irreversibility Comprised of 10 chapters this volume begins with an introduction to the fundamental nature of heat and of gases along with Boltzmann's work on the kinetic theory of gases and statistical mechanics Energy dissipation the statistical nature of the second law of thermodynamics and the eternal return and the recurrence paradox are also considered The first chapter examines the dynamical theory of gases and its application to the explanation of various properties of gases the known chemical relation between the density of a gas and its equivalent weight commonly called the Law of Equivalent Volumes and the diffusion of one gas through another Subsequent chapters focus on the thermal equilibrium of gas molecules the three body problem and the equations of dynamics the mechanical theory of heat and the relation of a general mechanical theorem to the second law of thermodynamics A mechanical explanation of irreversible processes is also offered This book will be useful for students of physics at the advanced undergraduate or beginning postgraduate level **Foundations of the Molecular Theory** John Dalton, Amedeo Avogadro, Joseph Louis Gay-Lussac, 1911 *Kinetic Molecular Theory* Charles Proteus Steinmetz, American Electrochemical Society. Meeting, 1908* Topics in Kinetic Theory Thierry Passot, Catherine Sulem, P. L. Sulem, 2005 This book covers a variety of topics related to kinetic theory in neutral gases and magnetized plasmas with extensions to other systems such as quantum plasmas and granular flows A comprehensive presentation is given for the Boltzmann equations and other kinetic equations for a neutral

gas together with the derivations of compressible and incompressible fluid dynamical systems and their rigorous justification. Several contributions are devoted to collisionless magnetized plasmas. Rigorous results concerning the well posedness of the Vlasov-Maxwell system are presented. Special interest is devoted to asymptotic regimes where the scales of variation of the electromagnetic field are clearly separated from those associated with the gyromotion of the particles. This volume collects lectures given at the Short Course and Workshop on Kinetic Theory organized at the Fields Institute of Mathematical Sciences in Toronto during the Spring of 2004. *Kinetic Theory*, 1921. **Kinetic Theory** J.M. Pendlebury, 1985-09-01. The theory of the motions of atoms and molecules in a fluid provides the key to understanding thermodynamics yet many students find difficulty in evaluating and using the basic expressions of kinetic theory. This book is designed to lead such students through the fundamental assumptions underlying the theory and to build up a clear physical understanding of the distribution of molecular speeds within a gas. The bulk of the text outlines the relationship between molecular motions and the macroscopic properties of gases including the equation of state for a perfect gas and for a van der Waals gas and derives a variety of transport coefficients and thermodynamics relations. **Applications of the Kinetic Theory** William Pingry Boynton, 2015-06-25. Excerpt from *Applications of the Kinetic Theory To Gases Vapors Pure Liquids and the Theory of Solutions*. This book presupposes a moderate acquaintance with the fundamentals of physics and chemistry and a mathematical equipment involving familiarity with the differential calculus and at least the notation of the integral calculus. It embodies a course of lectures given at the University of California during the academic years 1898-1901 but includes for the sake of greater completeness of treatment some matter not given in the lectures. For detailed information regarding many of the topics mentioned as for instance osmotic pressure and electrolysis and for a statement of the experimental basis for the theories presented the reader is referred to the standard texts which treat of these topics in extenso. The intent in this volume is not so much to discuss these facts and theories by themselves as to present their possible or probable relations to each other in the light of the kinetic theory. About the Publisher: Forgotten Books publishes hundreds of thousands of rare and classic books. Find more at www.forgottenbooks.com. This book is a reproduction of an important historical work. Forgotten Books uses state-of-the-art technology to digitally reconstruct the work, preserving the original format whilst repairing imperfections present in the aged copy. In rare cases, an imperfection in the original, such as a blemish or missing page, may be replicated in our edition. We do, however, repair the vast majority of imperfections successfully; any imperfections that remain are intentionally left to preserve the state of such historical works. **Foundations of the Molecular Theory** John Dalton, Amedeo Avogadro, 2017-07-25. From the Preface: THE papers here reprinted in chronological order serve to exhibit the historical development of the idea of a connection existing between the number of particles in different gases and the volume they occupy. It will be seen that Dalton first entertains the notion that equal volumes of different gases may contain the same number of ultimate particles at equal temperature and pressure but that he is legitimately forced to reject this

assumption conceiving no distinction between the atom and the molecule of an element Gay Lussac's important experimental work on the combining volumes of gases then shows the necessity of a simple relation between the ultimate particles of gases and their volumes although he does not point this out in his paper Dalton however perceives the necessity and characteristically concludes by doubting the accuracy of Gay Lussac's experiments Avogadro finally accepts both Dalton's theory and Gay Lussac's data and teaches how to reconcile them by distinguishing between the atom and the molecule of an elementary gas It has not been thought necessary to reprint the letter of Ampere to Berthollet Annales de Chimie 90 43 86 1814 since that paper contains no advance on the views of Avogadro published three years earlier its author simply drawing the same conclusions from the same premises The English version of the French originals will probably be found more faithful than elegant especially so in the case of Avogadro's paper where the French is always clumsy and occasionally obscure

Kinetic Theory of Bimolecular Chemical Reaction, Diffusive Drag, and Other Processes in a Gas Mixture Richard A. Hord, 1967 For two gas species with a temperature difference and a bulk velocity difference neither of which is necessarily small in magnitude the kinetic theory of gases has been used to derive explicit expressions for the collision frequency diffusive drag force molecular translational energy transfer rate and the bimolecular chemical reaction frequency The derivations which are based upon hypothesized mutual collision diameters activation energies and steric factors are of interest in connection with theoretical studies of low pressure gas mixtures with large departures from equilibrium A binary temperature concept is introduced as an aid in condensing and interpreting the expressions derived from the kinetic theory The expression derived for the diffusive drag force is used to give a more definite form to the equations of motion of the individual species in a mixture of several interdiffusing gases Kinetic Processes in Gases and Plasmas A Hochstim, 2012-12-02 Kinetic Processes in Gases and Plasmas provides a survey of studies on transport and chemical kinetic processes in high temperature gases and plasmas The book is concerned with conditions produced by the interaction of an object with the atmosphere at hypersonic velocities The text also provides a foundation for the flow field equations which include chemical reactions and other transport processes and to present in some detail the microscopic considerations underlying these calculations Chapters are devoted to the discussion of topics such as the molecular theory of transport equations transport processes in ionized gases and inelastic energy transfer processes and chemical kinetics Aerospace engineers physicists chemists and astrophysicists will find the book a good reference material An Introduction to Molecular Kinetic Theory Joel Henry Hildebrand, 1963 **Molecular Theory of Gases and Liquids** Joseph Oakland Hirschfelder, 2003-01 Foundations of the Molecular Theory, 1902 **Foundations of the Molecular Theory** John Dalton, Joseph Louis Gay-Lussac, Amedeo Avogadro, 1893 Foundations of Molecular Theory, Comprising Papers and Extracts John Dalton, Joseph Louis Gay-Lussac, Amedeo Avogadro, 1961 **Foundations of the Molecular Theory**, 1893 The Molecular Theory of Gases and Liquids Joseph O. Hirschfelder, Joseph Oakland Hirschfelder, Charles Francis

Curtiss, Robert Byron Bird, 1954 An essential cross disciplinary reference for molecular interactions Molecular Theory of Gases and Liquids offers a rigorous comprehensive treatment of molecular characteristics and behaviors in the gaseous and fluid states A unique cross disciplinary approach provides useful insight for students of chemistry chemical engineering fluid dynamics and a variety of related fields with thorough derivations and in depth explanations throughout Appropriate for graduate students and working scientists alike this book details advanced concepts without sacrificing depth of coverage or technical detail

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