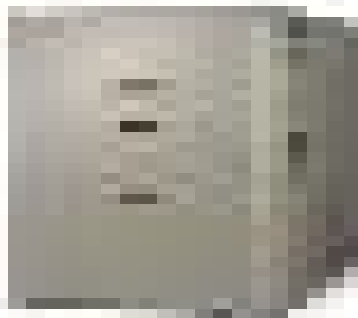


LABORATORY

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INTRODUCTION
The purpose of this experiment is to determine the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction is exothermic and produces oxygen gas and water. The rate of reaction is measured by the volume of oxygen gas evolved over a fixed period of time. The experiment is carried out at different temperatures and the results are plotted on a graph of rate of reaction versus temperature. The graph shows that the rate of reaction increases with increasing temperature. This is due to the fact that at higher temperatures, the molecules have more kinetic energy and are therefore more likely to collide and react. The experiment also demonstrates the effect of a catalyst on the rate of reaction. Potassium iodide acts as a catalyst for the reaction, and the rate of reaction is found to be much faster in the presence of the catalyst than in its absence. The experiment is a classic demonstration of the effect of temperature and catalysts on the rate of reaction, and it is a useful tool for teaching students about chemical kinetics.

Luxaire Oil Furnace Manual

WJ Hussar



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