

Integrated Science I

Chapter 6 Section 1 & 2 Quiz Review

1. Give a real world example of a compound.
Table salt, baking soda
2. What holds atoms and ions together?
A chemical bond
3. Identify what the following types of bond models are used for:
 - a. Ball-and-stick – *shows bond angle and bond length*
 - b. Space-filling – *shows space occupied by atoms*
 - c. Structural Formula – *shows symbols of elements used*
4. Strong bonds cause atoms to have what characteristics?
Strong attraction causes high melting and boiling points of compounds
5. How does the attraction between molecules compare to the attraction between ions?
The attraction between molecules is weaker. The bond between hydrogen atoms is weaker than the bond between the hydrogen molecule and oxygen atom in water.
6. Salt has a higher melting point than sugar because of what characteristic?
Salt is formed by a ionic bond and sugar is formed by a covalent bond
7. Why would a compound with only two atoms not have a bond angle?
Bond angles form when there is two or more bonds to the same atom.
8. Why do atoms form bonds? *Bonds are formed to create compounds and to stabilize atoms that do not have a full valence electron level.*
9. Identify what types of atoms make the following types of bonds:
 - a. Metallic metal + metal
 - b. Ionic metal + non-metal
 - c. Covalent non-metal + non-metal
10. What are covalent bonds and how do these bonds affect the boiling point of a substance?
Covalent bonds form when atoms share electrons. This type of bond causes the boiling and melting points of substances to be lower.
11. What are ionic bonds and how do these bonds affect the boiling point of a substance?
Ionic bonds form when atoms gain or lose electrons. This type of bond causes the boiling and melting points of substances to be higher.
12. What affect do metallic bonds have on the properties of metals? *Metals are able to be more flexible since the electrons and atoms can move more freely. It gives metals properties such as being malleable and ductile.*

Integrated Science Chapter 19 Review Answers

Yicheng Fang



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flower orientation the flowers point upward or spread or curve outward flower petal color pink to red white flower petal length 3 4 5 mm flower shape the flower has an urn shaped corolla constricted at the opening the flower is bell shaped

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general information bulb solitary or clustered applanate globose to cylindric ovoid tunic purple red brown red pale brown red or yellow to pale yellow papery to thinly leathery entire leaves shorter than scape 0 5 2 cm wide terete fistulose

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stem underground bulb leaf a cluster of radical leaves emerges from the underground bulb cylindrical and fleshy having

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the floral formula of allium cepa onion of the liliaceae family is as follows here the symbols represent systematic position of
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the flowers of allium are described as bowl shaped flowers in morphological terms with hidden nectarines generally they are
in under part of the ovary onion flowers are a good source of floral nectar and pollen the nectar produced from the nectaries
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allium cepa is an evergreen bulb growing to 0.6 m 2ft see above for usda hardiness it is hardy to uk zone 5 and is not frost tender it is in flower from june to july the species is hermaphrodite has both male and female organs and is pollinated by bees insects suitable for light sandy and medium loamy soils and prefers well drained soil

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allium is a genus of monocotyledonous flowering plants with hundreds of species including the cultivated onion garlic scallion shallot leek and chives the generic name allium is the latin word for garlic and the type species for the genus is

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