# ACIDS IN CHEMISTRY

Acids dissolve in water ( $H_2O$ ) and form ions. The following table consists of a list of ternary acids. Complete the anion and the cation columns and answer the given questions.

Acid	Name of the Acid in Aqueous Solution	Cation	Anion	Polyatomic Anion Name
нсю,	Chloric acid	H³O.		
H <sub>2</sub> SO <sub>3</sub>	Sulfurous acid			
H <sub>2</sub> SO <sub>4</sub>	Sulfuric acid			Sulfate
H <sub>3</sub> PO <sub>3</sub>	Phosphorous acid	3H3O-		Phosphite
H <sub>3</sub> PO <sub>4</sub>	Phosphoric acid		PO <sub>4</sub> 3-	
HNO <sub>3</sub>	Nitric acid			
HNO <sub>2</sub>	Nitrous acid		NO <sup>2-</sup>	Nitrite
н,со,	Carbonic acid			

a) How do ternary acids differ from binary acids in their structure?

b) What number does the prefix "ter" refer to? \_\_\_\_\_\_

c) When the polyatomic anion name ends with "-ate", the acid name ending is \_\_\_\_\_\_.

d) When the polyatomic anion name ends with "-ite", the acid name ending is \_\_\_\_\_\_ .

e) If the prefix "hydro-" were used to name a ternary acid, what problem would it cause to name HCIO<sub>3</sub>?

f) Write a rule for naming ternary acid.

g) Predict the formula for chlorous acid \_\_\_\_\_\_.

# **Instructional Fair Inc Chemistry If8766 Naming Acids**

**Yi-Tong Ma** 

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