

1. $\underline{\quad\quad}\text{Zn} + \underline{2}\text{HCl} \rightarrow \underline{\quad\quad}\text{ZnCl}_2 + \underline{\quad\quad}\text{H}_2$
2. $\underline{4}\text{Fe} + \underline{13}\text{O}_2 \rightarrow \underline{2}\text{Fe}_2\text{O}_3$
3. $\underline{4}\text{P} + \underline{5}\text{O}_2 \rightarrow \underline{\quad\quad}\text{P}_4\text{O}_{10}$
4. $\underline{\quad\quad}\text{N}_2 + \underline{3}\text{H}_2 \rightarrow \underline{2}\text{NH}_3$
5. $\underline{\quad\quad}\text{PbO}_2 \rightarrow \underline{\quad\quad}\text{Pb} + \underline{\quad\quad}\text{O}_2$
6. $\underline{2}\text{C}_6\text{H}_6 + \underline{15}\text{O}_2 \rightarrow \underline{12}\text{CO}_2 + \underline{6}\text{H}_2\text{O}$
7. $\underline{\quad\quad}\text{C}_2\text{H}_5\text{OH} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO}_2 + \underline{3}\text{H}_2\text{O}$
8. $\underline{3}\text{Fe} + \underline{4}\text{H}_2\text{O} \rightarrow \underline{\quad\quad}\text{Fe}_3\text{O}_4 + \underline{4}\text{H}_2$
9. $\underline{\quad\quad}\text{SiO}_2 + \underline{4}\text{HF} \rightarrow \underline{\quad\quad}\text{SiF}_4 + \underline{2}\text{H}_2\text{O}$
10. $\underline{3}\text{CO} + \underline{\quad\quad}\text{FeO}_4 \rightarrow \underline{\quad\quad}\text{FeO} + \underline{3}\text{CO}_2$
11. $\underline{\quad\quad}\text{NH}_4\text{NO}_2 \rightarrow \underline{\quad\quad}\text{N}_2 + \underline{2}\text{H}_2\text{O}$
12. $\underline{\quad\quad}\text{N}_2\text{H}_4 + \underline{\quad\quad}\text{O}_2 \rightarrow \underline{\quad\quad}\text{N}_2 + \underline{2}\text{H}_2\text{O}$
13. $\underline{3}\text{NO}_2 + \underline{\quad\quad}\text{H}_2\text{O} \rightarrow \underline{2}\text{HNO}_3 + \underline{\quad\quad}\text{NO}$
14. $\underline{2}\text{KClO}_3 \rightarrow \underline{2}\text{KCl} + \underline{3}\text{O}_2$
15. $\underline{\quad\quad}\text{C}_{12}\text{H}_{22}\text{O}_{11} + \underline{12}\text{O}_2 \rightarrow \underline{12}\text{CO}_2 + \underline{11}\text{H}_2\text{O}$
16. $\underline{4}\text{NH}_3 + \underline{3}\text{O}_2 \rightarrow \underline{6}\text{H}_2\text{O} + \underline{2}\text{N}_2$
17. $\underline{\quad\quad}\text{CH}_3\text{COOH} + \underline{\quad\quad}\text{O}_2 \rightarrow \underline{2}\text{CO}_2 + \underline{2}\text{H}_2\text{O}$
18. $\underline{2}\text{C}_2\text{H}_6 + \underline{7}\text{O}_2 \rightarrow \underline{4}\text{CO}_2 + \underline{6}\text{H}_2\text{O}$
19. $\underline{\quad\quad}\text{C}_6\text{H}_8 + \underline{10}\text{O}_2 \rightarrow \underline{8}\text{CO}_2 + \underline{4}\text{H}_2\text{O}$
20. $\underline{2}\text{C}_4\text{H}_{10} + \underline{13}\text{O}_2 \rightarrow \underline{8}\text{CO}_2 + \underline{10}\text{H}_2\text{O}$
21. $\underline{\quad\quad}\text{N}_2 + \underline{3}\text{H}_2 \rightarrow \underline{2}\text{NH}_3$
22. $\underline{3}\text{Fe} + \underline{4}\text{H}_2\text{O} \rightarrow \underline{\quad\quad}\text{Fe}_3\text{O}_4 + \underline{4}\text{H}_2$
23. $\underline{3}\text{Fe} + \underline{2}\text{O}_2 \rightarrow \underline{\quad\quad}\text{Fe}_3\text{O}_4$
24. $\underline{3}\text{NH}_4\text{OH} + \underline{\quad\quad}\text{AlCl}_3 \rightarrow \underline{\quad\quad}\text{Al(OH)}_3 + \underline{3}\text{NH}_4\text{Cl}$
25. $\underline{\quad\quad}\text{Na}_2\text{CO}_3 + \underline{\quad\quad}\text{H}_2\text{O} + \underline{\quad\quad}\text{CO}_2 \rightarrow \underline{2}\text{NaHCO}_3$
26. $\underline{\quad\quad}\text{MnO}_2 + \underline{4}\text{HCl} \rightarrow \underline{\quad\quad}\text{MnCl}_2 + \underline{2}\text{H}_2\text{O} + \underline{\quad\quad}\text{Cl}_2$
27. $\underline{\quad\quad}\text{PBr}_3 + \underline{3}\text{H}_2\text{O} \rightarrow \underline{\quad\quad}\text{H}_3\text{PO}_3 + \underline{3}\text{HBr}$
28. $\underline{2}\text{ZnS} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{ZnO} + \underline{\quad\quad}\text{SO}_2$
29. $\underline{2}\text{NH}_3 + \underline{3}\text{O}_2 \rightarrow \underline{\quad\quad}\text{N}_2 + \underline{6}\text{H}_2\text{O}$
30. $\underline{\quad\quad}\text{Na}_2\text{B}_4\text{O}_7 + \underline{\quad\quad}\text{H}_2\text{SO}_4 + \underline{5}\text{H}_2\text{O} \rightarrow \underline{\quad\quad}\text{Na}_2\text{SO}_4 + \underline{4}\text{H}_3\text{BO}_3$
31. $\underline{2}\text{NH}_3 + \underline{2}\text{O}_2 \rightarrow \underline{\quad\quad}\text{NO} + \underline{3}\text{H}_2\text{O}$
32. $\underline{4}\text{HNO}_3 \rightarrow \underline{2}\text{H}_2\text{O} + \underline{4}\text{NO}_2 + \underline{\quad\quad}\text{O}_2$
33. $\underline{4}\text{HCl} + \underline{2}\text{HNO}_3 \rightarrow \underline{2}\text{Cl}_2 + \underline{2}\text{NOCl} + \underline{3}\text{H}_2\text{O}$
34. $\underline{2}\text{Ca}_3(\text{PO}_4)_2 + \underline{6}\text{SiO}_2 \rightarrow \underline{6}\text{CaSiO}_3 + \underline{\quad\quad}\text{P}_4\text{O}_{10}$
35. $\underline{\quad\quad}\text{C}_6\text{H}_{12}\text{O}_6 \rightarrow \underline{2}\text{C}_2\text{H}_5\text{OH} + \underline{2}\text{CO}_2$
36. $\underline{2}\text{C}_8\text{H}_{18} + \underline{25}\text{O}_2 \rightarrow \underline{16}\text{CO}_2 + \underline{18}\text{H}_2\text{O}$
37. $\underline{2}\text{CH}_4 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO} + \underline{4}\text{H}_2\text{O}$
38. $\underline{2}\text{C}_2\text{H}_2 + \underline{3}\text{O}_2 \rightarrow \underline{4}\text{CO}_2 + \underline{2}\text{H}_2\text{O}$
39. $\underline{6}\text{H}_2\text{O} + \underline{6}\text{CO}_2 \rightarrow \underline{\quad\quad}\text{C}_6\text{H}_{12}\text{O}_6 + \underline{6}\text{O}_2$
40. $\underline{\quad\quad}\text{Cu} + \underline{4}\text{HNO}_3 \rightarrow \underline{\quad\quad}\text{Cu(NO}_3)_2 + \underline{2}\text{NO} + \underline{2}\text{H}_2\text{O}$

Instructional Fair Balancing Chemical Equations

Robert R. Redfield



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science it is a scientific discipline that involves the study of numbers quantities and shapes and their relationships and properties

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