

সুতরাং, বিশেষ MY_2 অক্সিজেন হবে।



↑↑↑↑ ivlv fv+jv:

কিছু ভ্রাম্যমাণের কোনো প্রকণে একটি লবণ প্রস্তুত থাকবে যদি অক্সিজেন হবে তা নির্ভর করে লবণটির প্রকৃতিতে তাপমাত্রার উপর। যদি প্রকণে কোনো লবণের উপস্থিতিতে অক্সিজেনের আয়নিক তাপমাত্রার মান লবণটির প্রকৃতিতে তাপমাত্রার মানকে অতিক্রম করতে না পারে তবে লবণটি প্রকণে প্রস্তুত অক্সিজেন থেকে। আর যদি লবণের আয়নিক তাপমাত্রার মান (K_{sp}) লবণটির প্রকৃতিতে তাপমাত্রার মান (K_{sp}) কে অতিক্রম করে তবে লবণটি প্রকণ থেকে অক্সিজেন হয়।

অর্থাৎ,

$K_{sp} < K_{sp}$: প্রকণটি অসম্পূর্ণ প্রকণ

$K_{sp} = K_{sp}$: প্রকণটি সম্পূর্ণ প্রকণ

$K_{sp} > K_{sp}$: প্রকণ থেকে প্রব অক্সিজেন হয়

III A পাত্রে PN প্রকণ যোগ করলে তা নিম্নলিখিত অক্সিজেন হয়।

PN Hooooo P⁺ + N⁻

প্রকণে, PN প্রকণের ঘনমাত্রা, 0.1M হবে-

P⁺ আয়নের ঘনমাত্রা, [P⁺] = 0.1 M

N⁻ আয়নের ঘনমাত্রা, [N⁻]₁ = 0.1 M

এখানে, PN প্রকণে N⁻ আয়ন উপস্থিত। অন্যদিকে A পাত্রে MN₃ লবণের একটি অক্সিজেন N⁻ উপস্থিত থাকবে A পাত্রে PN প্রকণ যোগ করলে সমস্ত অক্সিজেন প্রকণের কারণে MN₃ এর প্রকৃতিতে প্রব হবে।

যদি, MN₃ এর লবণ প্রকৃতি, S

প্রকণে MN₃ নিম্নলিখিত অক্সিজেন হয়,

MN₃ Hooooo M³⁺ + 3N⁻

সুতরাং, M³⁺ আয়নের ঘনমাত্রা, [M³⁺] = S

N⁻ আয়নের ঘনমাত্রা, [N⁻]₂ = 3S

সেহেতু, N⁻ আয়ন PN ও MN₃ উভয় স্ট্রিং উপস্থিত, সেহেতু

N⁻ আয়নের মোট ঘনমাত্রা হবে, [N⁻] = (3S + 0.1)M

এখন MN₃ এর আয়নিক তাপমাত্রা,

$$K_{sp} = [M^{3+}] \times [N^-]^3$$

যা, $2.7 \times 10^{-11} = S \times (3S + 0.1)^3$ [∵ 'খ' হচ্ছে, $K_{sp} = 2.7 \times 10^{-11}$]

$$= S \times [(3S)^3 + 3 \times (3S)^2 \times 0.1 + 3 \times 3S \times (0.1)^2 + (0.1)^3]$$

$$= S \times [27S^3 + 2.7S^2 + 0.09S + 0.001]$$

$$= 27S^4 + 2.7S^3 + 0.09S^2 + 0.001S$$

যা, $2.7 \times 10^{-11} = 0.001S$ [∵ S^4 , S^3 , S^2 এর মান খুবই নগণ্য]

$$\therefore S = \frac{2.7 \times 10^{-11}}{0.001}$$

$$= 2.7 \times 10^{-8} \text{ mol L}^{-1}$$

অর্থাৎ, MN₃ স্ট্রিংয়ের লবণ প্রকৃতি, $2.7 \times 10^{-8} \text{ mol L}^{-1}$

অতএব, A পাত্রে PN প্রকণ যোগ করলে MN₃ এর প্রকৃতি $10^{-8} \text{ mol L}^{-1}$ থেকে A পাত্রে $2.7 \times 10^{-8} \text{ mol L}^{-1}$ হবে।

Hsc Chemistry Note Bd

S Marginson



Hsc Chemistry Note Bd:

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