



Hp 1200 Manual

**United States. Department of the
Army**



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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4$
 $3\text{CH}_4 + 2\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
 $8\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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web 1 78 348 stoichiometry helps us determine how much substance is needed or is present things that can be measured are reactants and products mass molecular weight chemical equations formulas stoichiometric coefficient the stoichiometric coefficient or stoichiometric number is the number of molecules that participate in the reaction

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $2\text{C}_6\text{H}_{10}\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + \text{GaBr} + \text{NaBr} + \text{SnO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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