

Qualitative analysis of an organic compound showed that it contained only carbon, hydrogen and oxygen. A quantitative study of the same compound was performed, in which a 0.5096 g sample was burnt in excess oxygen to produce 0.4160 g of water and 700.7 mL of carbon dioxide, collected at 100.0 °C and 102.8 kPa.

373K

(a) Determine the empirical formula of the compound.

(10 marks)

C	H	O
<p>0.7007 L of <math>\text{CO}_2</math></p> <p><math>n(\text{C}) = n(\text{CO}_2)</math></p> <p><math>PV = nRT \rightarrow n = \frac{PV}{RT}</math></p> <p><math>n(\text{CO}_2) = \frac{102.8 \times 0.7007}{8.315 \times 373}</math></p> <p><math>= 2.3218 \times 10^{-2} \text{ mol}</math></p> <p><math>n(\text{C}) = nM = 0.27885 \text{ g}</math></p>	<p>0.416 g of <math>\text{H}_2\text{O}</math></p> <p><math>n = \frac{m}{M} = \frac{0.416}{18.016} = 2.3091 \times 10^{-2} \text{ mol}</math></p>	<p><math>m(\text{O}) = 0.5096 - m(\text{C}) - m(\text{H})</math></p> <p><math>= 0.18423 \text{ g}</math></p> <p><math>n = \frac{m}{M} = \frac{0.18423}{16} = 1.1514 \times 10^{-2}</math></p>

EMPIRICAL FORMULA

EXAM QUESTION

$$\frac{2.3218}{1.1514} = 2$$

# Empirical Formula Gravimetric Questions Answers

**Michael R. Lindeburg**



## **Empirical Formula Gravimetric Questions Answers:**

*Chemical Calculations with Explanatory Notes, Problems, and Answers, Specially Adapted for Use in Colleges and Science Schools* Richard Lloyd Whiteley, 1896      *Chemistry: for B.Sc. Students Semester II (NEP-2020 KSHCE)* R L Madan, This textbook has been conceptualised for universities of Karnataka as per the recommended National Education Policy NEP 2020 to meet the needs of B Sc students of Chemistry Maintaining the traditional approach to the subject this textbook lucidly explains the basics of Inorganic and Physical Chemistry Important topics such as atomic structure periodicity of elements kinetic theory of gases liquid state liquid crystals and types of solids are aptly discussed to give an overview of inorganic and physical chemistry Laboratory work 33 experiments has also been included to help students achieve sound conceptual understanding and learn experimental procedures      **CliffsAP 5 Chemistry Practice Exams** Gary S

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*Environmental Science Practical* R.K. Kamble, 2018-08-01 This book is primarily prepared to cater students of undergraduate postgraduate research scholars and faculty members in Environmental Science Environmental Engineering Environmental Technology of universities institutes of India and abroad It provides sufficient theoretical and practical knowledge about

various environmental parameters so as to have a clear understanding of them The book comprises of four parts viz air water soil and noise Each part further contains various parameters involved in them except noise Number of questions and answers on each parameter are presented in lucid and concise manner so as to make all the aspects of it understandable In addition to this a number of appendixes are also appended which will provide additional knowledge on these parameters for overall understanding of them

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these updates Sections discuss pollutant characterization pollutant transformation and environmental partitioning Air partitioning physical transport of air pollutants air pollution biogeochemistry and thermal reactions are also thoroughly explored The author then carefully examines air pollution risk calculations control technologies and dispersion models The text wraps with discussions of economics and project management reliability and failure and air pollution decision making Provides real life current cases as examples of quantitation of emerging air pollution problems Includes straightforward derivation of equations giving practitioners and instructors a direct link between first principles of science and applications of technologies Presents example calculations that make scientific theory real for the student and practitioner *Druggists' Circular*, 1912

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*Nationalizing Science* Alan J.

Rocke, 2000-11-08 After looking at the early careers of Wurtz's two mentors Liebig and Jean Baptiste Dumas Rocke describes Wurtz's life and career in the politically complex period leading up to 1853. He then discusses the turning point in Wurtz's intellectual life, his conversion to the reformed chemistry of Laurent Gerhardt and Williamson, and his efforts to persuade his colleagues of the advantages of the new system. In 1869 Adolphe Wurtz (1817-1884) called chemistry a French science. In fact, however, Wurtz was the most internationalist of French chemists. Born in Strasbourg and educated partly in the laboratory of the great Justus Liebig, he spent his career in Paris where he devoted himself to introducing German ideas into French scientific circles. His life therefore provides an excellent vehicle for considering the divergent trajectories of French and German chemistry and by extension French and German science during this crucial period. After looking at the early careers of Wurtz's two mentors Liebig and Jean Baptiste Dumas, Rocke describes Wurtz's life and career in the politically complex period leading up to 1853. He then discusses the turning point in Wurtz's intellectual life, his conversion to the reformed chemistry of Laurent Gerhardt and Williamson, and his efforts, social and political as well as scientific, to persuade his colleagues of the advantages of the new system. He looks at political patronage or the lack thereof and at the insufficient material support from the French government during the middle decades of the century. From there, Rocke goes on to examine the rivalry between Wurtz and Marcellin Berthelot, the debate over atoms versus equivalents, and the reasons for Wurtz's failure to win acceptance for his ideas. The story offers insights into the changing status of science in this period and helps to explain the eventual course of both French and German chemistry.

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