

Chapter 2 Study Guide

CHAPTER 2 The Chemical Basis of Life I: Atoms, Molecules, and Water

- Describe the location, charge and mass of protons, neutrons and electrons, within an atom.**
 - Neutrons and Protons are densely packed in the Nucleus.
 - Electrons move about the Nucleus in Orbitals.
 - Neutrons and Protons have a mass of one.
- Be able to calculate the Atomic Weight of an atom/element, if the number of protons, electrons and neutrons is known.**
 - Atomic Mass Number = Protons + Neutrons
- Be able to calculate the Atomic Number of an atom/element, if the number of protons, electrons and neutrons is known.**
 - Atomic Number = Number of Protons
- Be able to write the symbols, atomic numbers, and mass numbers of the biologically important elements (C, H, O, & N).**
 - C → Carbon, 12 Protons, 12.011 Mass
 - H → Hydrogen, 1 Proton, 1.008 Mass
 - O → Oxygen, 8 Protons, 15.999 Mass
 - N → Nitrogen, 7 Protons, 14.007 Mass
- Be able to explain how isotopes of an element differ in terms of the number of electrons, neutrons and protons.**
 - Isotopes differ in the number of neutrons in the Nucleus.
 - The number of Protons does not change.
 - The number of Electrons does not change.
- Be able to correctly place any number of electrons in an atom's first two energy level shells.**
 - First Energy Level requires 2 Electrons.
 - Any Energy Levels after the first require 8 Electrons
- Differentiate between atoms, molecules and ions.**
 - Atoms are single neutral particles.
 - Molecules are neutral particles made of two or more atoms bonded together.
 - Ions are positively or negatively charged particles (atoms)
- Differentiate between molecules and compounds.**
 - All compounds are molecules, but not all molecules are compounds.
 - Molecules are made up of any two or more atoms bonded together, but Compounds are made up of two or more different atoms bonded together.
- What is a valence electron?**
 - Valence Electrons are Electrons located in the outermost Energy Shell of an atom. Since they are the farthest from the Nucleus, they can participate in bonds and reactions, therefore Valence Electrons determine an atom's reactivity, electronegativity, and the number of bonds it can form.
- Describe a covalent chemical bond, and what holds the two atoms together.**
 - A Covalent Bond is formed through the equal sharing of valence Electrons in two atoms, with the sharing process enabling each atom to fill its outer valence electrons.
- Differentiate between a polar and non-polar covalent bond.**
 - A Polar Covalent Bond results if the Electronegativity difference from a bond is between 0.5-1.6, meaning one end of the bond is partly negative while the other is partly positive.
 - A Non-Polar Covalent Bond results if the Electronegativity difference from a bond is between 0.2-0.5
- What is a Hydrogen Bond, and what holds the two atoms together?**
 - A Hydrogen Bond occurs only with Polar Molecules, as the partial opposite charges of the molecules will attract one another, usually placing a Hydrogen atom as the bridge between the molecules.
- What is an ionic bond, and what holds the two atoms together.**
 - An Ionic Bond is a bond that is formed from the transfer of Electrons from one atom to another, therefore making one element have more Electrons than Protons and the other have fewer Electrons than Protons.
 - Essentially, the more Electronegative atom strips Electrons from another atom to complete its outer-most valence shell.
- What is a van der Waals interaction (bond) and what holds the two atoms together?**

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Teacher's Guide to the Modern Biology Program James Howard Otto, Albert Towle, Elizabeth H. Crider, 1965

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