

Chemistry – Unit 4 Worksheet 4

DMC.5: *I can cite and interpret the evidence that supports the belief that some pure substances are made up of simpler particles (atoms) in a definite ratio.*

Answer the following questions. Be sure to show all mathematical work and reasoning and use complete sentences in explanations.

- Two compounds of hydrogen and oxygen are tested. Compound I contains 120.0 g of oxygen and 15.0 g of hydrogen. Compound II contains 32.0 g of oxygen and 2.0 g of hydrogen.
 - Determine the ratio of the mass of oxygen to the mass of hydrogen in each of the compounds.
 - Why are the compounds not the same?
 - What is significant about these mass ratios?
 - If compound I is water, what could be the formula of compound II?
- Nitrogen and oxygen combine to form a variety of compounds. The following data were collected for three different compounds of nitrogen and oxygen:

Analysis Data of Nitrogen & Oxygen Compounds	
Compound	Mass of Nitrogen that combines with every 1.0 g of Oxygen
A	1.750 g
B	0.8750 g
C	0.4375 g

- Additional evidence shows that the formula of compound B is NO. Sketch particle diagrams of molecules of all three compounds.
- Justify your representations above.
- There is no compound between Nitrogen and oxygen with a mass ratio of 2.00 grams of N for every 1.00 grams of oxygen. Why?

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