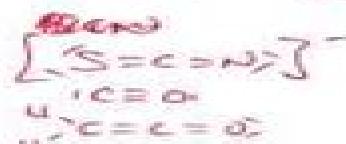


10 • Orbital Hybridization & Molecular Orbitals**PRACTICE TEST**

1. Which hybridization is associated with a steric number of 3?
 a) sp
 b) sp^2
 c) sp^3
 d) sp^3d
2. The molecule BrF_3 has a steric number of on the central atom?
 a) 3
 b) 4
 c) 5
 d) 6
3. What is the hybridization of Br in BrF_3 ?
 a) sp
 b) sp^2
 c) sp^3
 d) sp^3d
4. How many equivalent sp^3d orbitals are there?
 a) 3
 b) 1
 c) 5
 d) 6
5. What type of hybridization is associated with a square planar molecular shape?
 a) sp^2
 b) sp^3
 c) sp^3d^2
6. What shape for electron pairs is associated with sp^3d^2 hybridization?
 a) linear
 b) square planar
 c) bent
 d) tetrahedral
7. What hybridization is predicted for phosphorus in the PCl_3 molecule?
 a) sp^2
 b) sp^3
 c) sp
 d) sp^3d^2

tetrahedral shape - pyramidal

8. A double bond contains 1 sigma bond(s) and 2 pi bond(s).
 a) 0, 2
 b) 1, 2
 c) 2, 0
 d) 1, 1
9. What angle exists between orbitals in sp^2d^2 hybrid orbitals?
 a) 90.0°
 b) 180.0°
 c) 78.5°
 d) 120.0°
10. Which of the following elements is most likely to display sp^3d hybridization?
 a) oxygen
 b) nitrogen
 c) phosphorus
 d) carbon
 e) boron
11. How many sigma (σ) and pi (π) electrons pairs are in a carbon dioxide molecule?
 a) four σ and zero π
 b) three σ and two π
 c) one σ and three π
 d) two σ and two π
12. What is the hybridization of the oxygen atoms in CH_3OH and CO_2 , respectively?
 a) sp^3 , sp^3
 b) sp^3 , sp^2
 c) sp^2 , sp^2
13. All of the following species contain two π -bonds EXCEPT
 a) SCN^-
 b) CO
 c) H_3CCO



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